



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

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CHEMISTRY

0620/03

Paper 3 Theory (Core)

For Examination from 2016

SPECIMEN PAPER

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

A copy of the Periodic Table is printed on page 16.

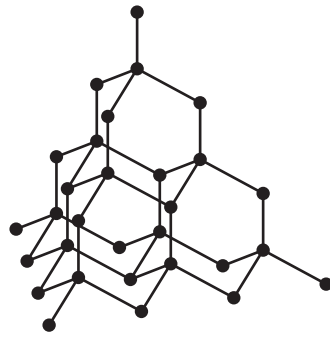
At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is accredited for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

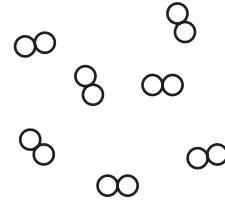
This document consists of **15** printed pages and **1** blank page.

1 The structures of diamond and chlorine are shown below.



diamond

● = carbon atom



chlorine

○ = chlorine atom

(a) Describe the structure of these two substances.
Use the list of words to help you.

covalent diatomic giant macromolecule molecule structure

diamond

.....

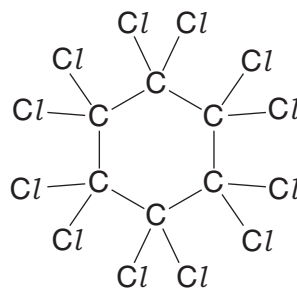
.....

chlorine

.....

..... [4]

(b) The structure of a compound containing carbon and chlorine is shown below.



What is the molecular formula of this compound?

..... [1]

(c) Chlorine is a halogen.

(i) State the colour of chlorine.

..... [1]

The table shows some properties of the halogens.

element	boiling point/°C	density in liquid state/g per cm ³	colour
fluorine	-188	1.51	yellow
chlorine	-35	1.56	
bromine	-7		red-brown
iodine	+114	4.93	grey-black

Use the information in the table to answer the following questions.

(ii) Predict the density of liquid bromine.

..... [1]

(iii) Describe the trend in boiling point of the halogens down the group.

..... [1]

(d) (i) Complete the word equation for the reaction of bromine with aqueous potassium iodide.

bromine + potassium iodide → +

..... [2]

(ii) Suggest why bromine does not react with aqueous potassium chloride.

..... [1]

(e) Potassium chloride is an ionic substance but iodine is a molecular substance. How do most ionic and molecular substances differ in their

solubility in water?

.....

electrical conductivity?

..... [2]

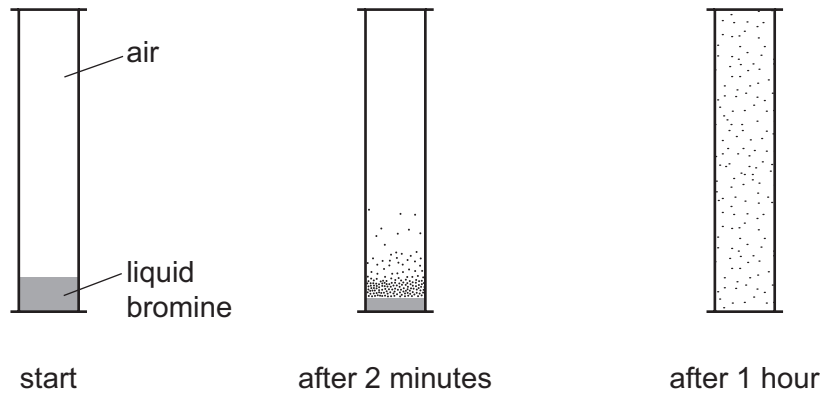
[Total: 13]

2 Bromine is an element in Group VII of the Periodic Table.

(a) State the formula for a molecule of bromine.

..... [1]

(b) A teacher placed a small amount of liquid bromine in the bottom of a sealed gas jar of air. After two minutes red-brown fumes were seen just above the liquid surface. After one hour the red-brown colour had spread completely throughout the gas jar.



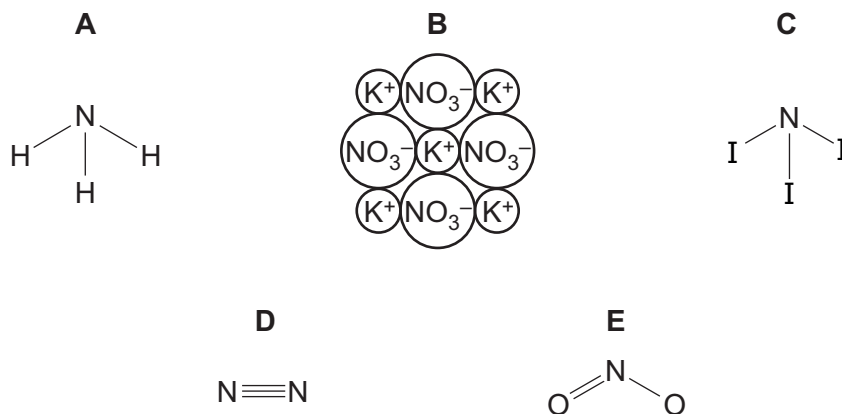
Use the kinetic particle model of matter to explain these observations.

.....

 [3]

[Total: 4]

3 The structures of some substances containing nitrogen are shown below.



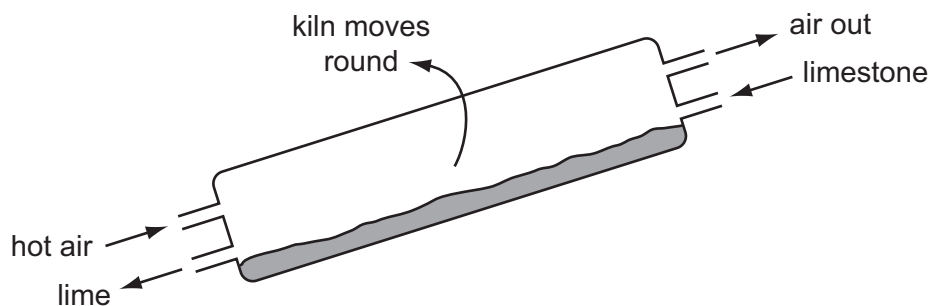
Answer the following questions by choosing from the structures **A**, **B**, **C**, **D** or **E**.
You can use each structure once, more than once or not at all.

Which structure represents

- | | | |
|---|---|-----|
| (a) an acidic oxide, | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |
| (b) an ionic structure, | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |
| (c) a gas which turns damp red litmus paper blue, | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |
| (d) a compound which is formed under conditions of high temperature and pressure in car engines, | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |
| (e) a molecule containing halogen atoms, | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |
| (f) a salt? | <input style="width: 40px; height: 30px;" type="checkbox"/> | [1] |

[Total: 6]

- 4 The diagram shows a rotary lime kiln used to make lime from limestone. Limestone is fed in at the top of the kiln and lime comes out at the bottom.



- (a) State the chemical name for lime?

..... [1]

- (b) State the name of the type of chemical reaction that takes place in the kiln.

..... [1]

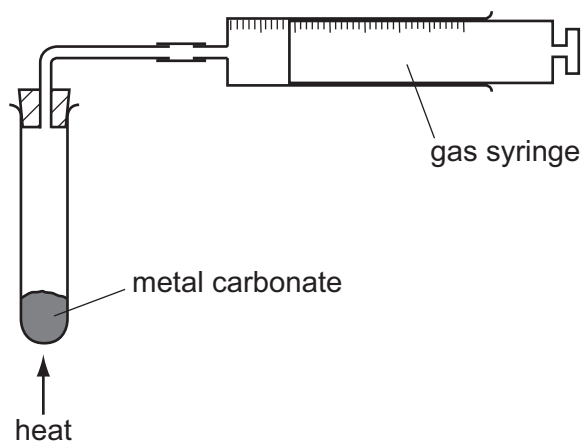
- (c) Suggest why the air coming out of the kiln has a greater percentage of carbon dioxide than the air entering the kiln.

..... [1]

- (d) State **one** use for lime.

..... [1]

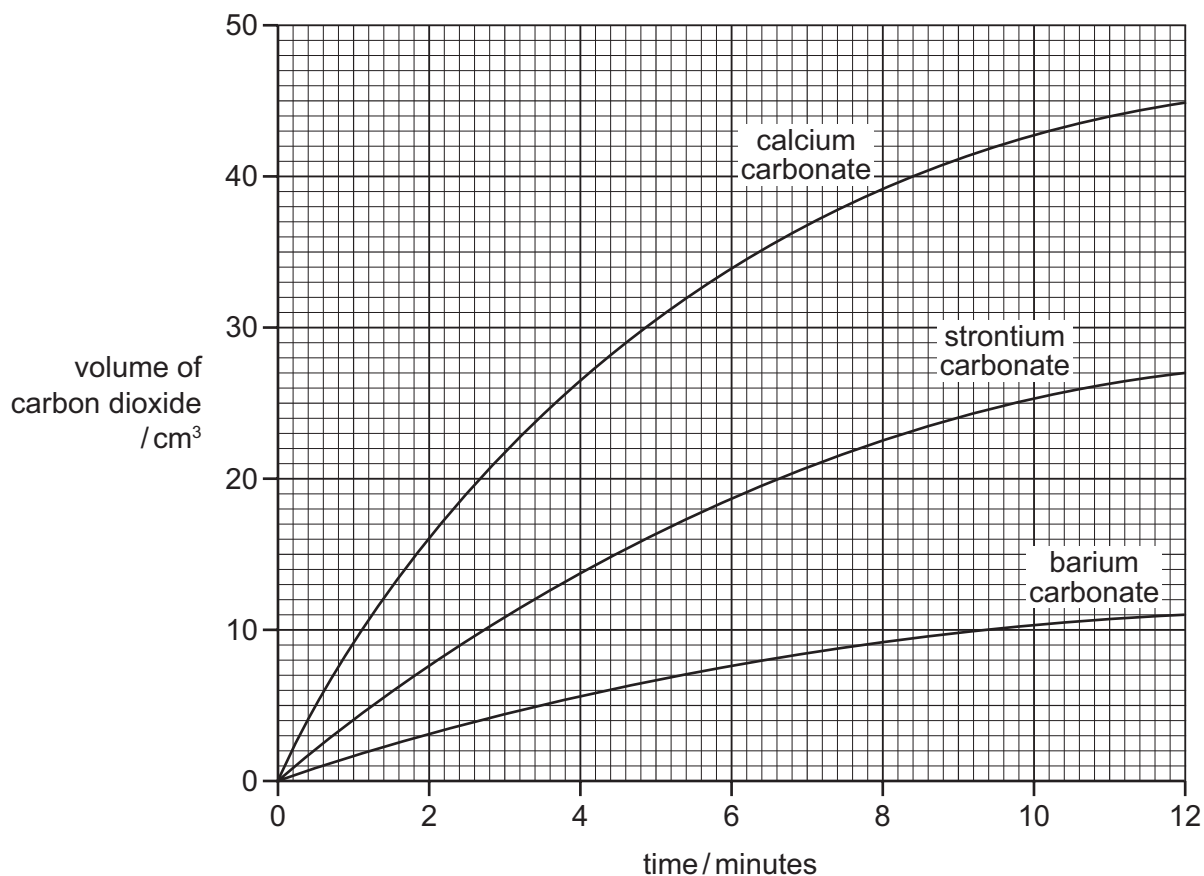
- (e) A student compared the rates of reaction of three metal carbonates. She measured the volume of gas released using the apparatus shown.



State **one** thing that must be kept constant if the rates of the three reactions are to be compared in a fair way.

..... [1]

- (f) The graph shows the volume of carbon dioxide released when the three metal carbonates were heated.



- (i) Which carbonate produced carbon dioxide at the highest rate?

..... [1]

- (ii) What volume of carbon dioxide was produced by strontium carbonate in twelve minutes?

..... [1]

- (iii) How do the rates of the reactions of these three metal carbonates relate to the position of calcium, strontium and barium in the Periodic Table?

.....

..... [2]

- (g) Describe how hydrochloric acid and limewater can be used to show that carbonate ions are present in calcium carbonate.

.....

.....

..... [3]

[Total: 12]

5 Iron is a transition element.

(a) State **three** properties of transition elements which are **not** shown by the Group I elements.

1.
2.
3. [3]

(b) The symbols for two isotopes of iron are shown below.



(i) How do these two isotopes differ in their atomic structure?

..... [1]

(ii) Determine the number of neutrons present in one atom of the isotope ${}_{26}^{57}\text{Fe}$.

..... [1]

(iii) Determine the number of electrons in one Fe^{3+} ion?

..... [1]

(c) Pure iron rusts very easily.

Describe and explain **one** method of preventing rusting.

method

explain why this method works

..... [2]

(d) Iron can be recycled.

Explain **two** advantages of recycling metals.

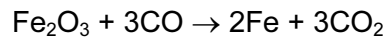
.....

.....

.....

..... [2]

- (e) In the blast furnace, iron(III) oxide reacts with carbon monoxide.



Which substance gets reduced in this reaction?
Explain your answer.

substance

explanation

..... [2]

- (f) (i) Carbon monoxide is a pollutant gas produced in motor car engines.
State why carbon monoxide is formed.

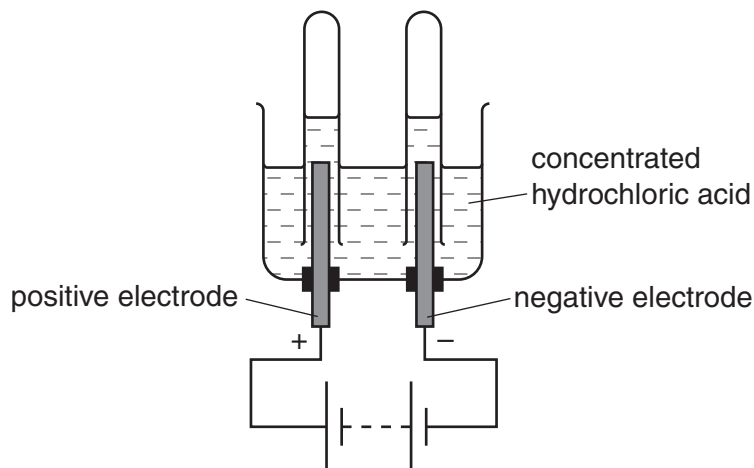
..... [1]

- (ii) State **one** harmful effect of carbon monoxide.

..... [1]

[Total: 14]

- 6 Concentrated hydrochloric acid can be electrolysed using the apparatus shown.



- (a) Define the term *electrolysis*?

.....
 [1]

- (b) What is the name given to the positive electrode?
 Put a ring around the correct answer.

anion **anode** **cathode** **cation** **electrolyte**

[1]

- (c) State the name of the gas given off at the negative electrode.

..... [1]

- (d) Complete the following sentence about electrolysis using words from the list.

inert **magnesium** **platinum** **reactive** **solid**

Electrodes made of graphite or are generally used in electrolysis
 because they are

[2]

(e) When concentrated hydrochloric acid is electrolysed, chlorine is released.

(i) Draw the shells and the electronic structure in an atom of chlorine.

[1]

(ii) Draw the electronic structure of a chlorine molecule.
Show only the outer electron shells.

[2]

(iii) Describe a test for chlorine.

test

result [2]

(f) Hydrochloric acid reacts with the base calcium hydroxide.

(i) Complete the word equation for this reaction.

hydrochloric acid + calcium hydroxide → +

..... [2]

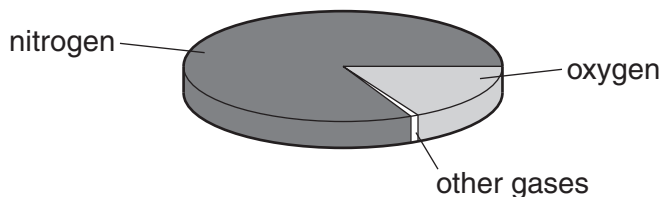
(ii) Hydrochloric acid also reacts with zinc.
Complete the symbol equation for this reaction.

$\text{Zn} + \dots\dots\text{HCl} \rightarrow \text{ZnCl}_2 + \dots\dots$

[2]

[Total: 14]

7 The pie chart shows the composition of air.



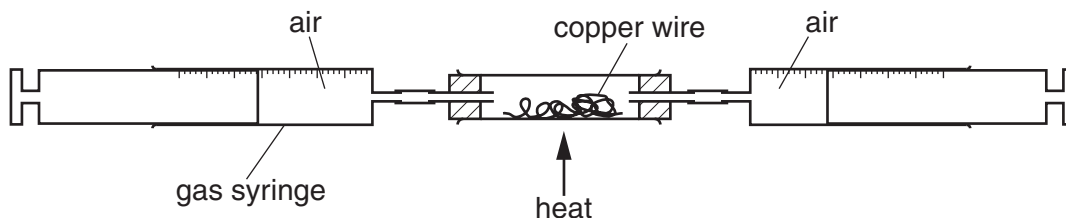
(a) (i) What is the percentage of nitrogen in the air?

..... [1]

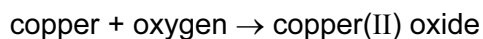
(ii) Apart from nitrogen and oxygen, state the names of **two** gases present in unpolluted air.

..... and [2]

(b) The percentage of oxygen in air can be found using the apparatus shown below.



Air is passed backwards and forwards over the heated copper using the syringes. The copper reacts with oxygen in the air.



As the experiment proceeds, suggest what happens to

(i) the total volume of air in the gas syringes,

..... [1]

(ii) the mass of the wire in the tube.

..... [1]

(c) State **one** use of copper.

..... [1]

[Total: 6]

8 Ethene, C₂H₄, is manufactured by cracking petroleum fractions.

(a) (i) What do you understand by the term *fraction*?

.....
 [1]

(ii) Complete the symbol equation for the manufacture of ethene from dodecane, C₁₂H₂₆.



(b) Two fractions obtained from the distillation of petroleum are refinery gas and gasoline. State **one** use of each of these fractions.

refinery gas

gasoline

[2]

(c) Ethene is an unsaturated hydrocarbon. What do you understand by the following terms?

unsaturated

hydrocarbon

[2]

(d) Ethene is used to make ethanol.

(i) Which of these reactions is used to make ethanol from ethene?
 Tick one box.

- catalytic addition of steam
- fermentation
- oxidation using oxygen
- reduction using hydrogen

[1]

(ii) Draw the structure of ethanol, showing all atoms and bonds.

[2]

- (e) Ethene is used to make poly(ethene).
Complete the following sentences about this reaction.
Use words from the list below.

additions carbohydrates catalysts monomers polymers

The ethene molecules which join to form poly(ethene) are the

The poly(ethene) molecules formed are [2]

[Total: 11]

Group																																		
I	II	III										IV	V	VI	VII	VIII																		
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 5px;"> 1 H hydrogen 1 </div> <div style="border: 1px solid black; padding: 5px; text-align: center;"> Key atomic number atomic symbol name relative atomic mass </div> </div>																																
3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21																
Li lithium 7	Be beryllium 9	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20	Na sodium 23	Mg magnesium 24	Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40	K potassium 39	Ca calcium 40	Sc scandium 45																
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55																
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 91	Nb niobium 93	Mo molybdenum 96	Tc technetium —	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106	Ag silver 108	Cd cadmium 112	In indium 115	Sn tin 117	Sb antimony 122	Te tellurium 128	I iodine 127	Xe xenon 131	55																
55	56	57–71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87																
Cs caesium 133	Ba barium 137	lanthanoids	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195	Au gold 197	Hg mercury 201	Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium —	At astatine —	Rn radon —	87																
87	88	89–103	104	105	106	107	108	109	110	111	112	114	114	116	116	—	—	—																
Fr francium —	Ra radium —	actinoids	Rf rutherfordium —	Db dubnium —	Sg seaborgium —	Bh bohrium —	Hs hassium —	Mt meitnerium —	Ds darmstadtium —	Rg roentgenium —	Cn copernicium —	Fl flerovium —	—	—	—	—	—	—																
<div style="display: flex; justify-content: space-between;"> <div style="width: 15%;"> lanthanoids 57 La lanthanum 139 </div> <div style="width: 15%;"> 60 Nd neodymium 144 </div> <div style="width: 15%;"> 61 Pm promethium — </div> <div style="width: 15%;"> 62 Sm samarium 150 </div> <div style="width: 15%;"> 63 Eu europium 152 </div> <div style="width: 15%;"> 64 Gd gadolinium 157 </div> <div style="width: 15%;"> 65 Tb terbium 159 </div> <div style="width: 15%;"> 66 Dy dysprosium 163 </div> <div style="width: 15%;"> 67 Ho holmium 165 </div> <div style="width: 15%;"> 68 Er erbium 167 </div> <div style="width: 15%;"> 69 Tm thulium 169 </div> <div style="width: 15%;"> 70 Yb ytterbium 173 </div> <div style="width: 15%;"> 71 Lu lutetium 175 </div> </div>																																		
<div style="display: flex; justify-content: space-between;"> <div style="width: 15%;"> actinoids 89 Ac actinium — </div> <div style="width: 15%;"> 90 Th thorium 232 </div> <div style="width: 15%;"> 91 Pa protactinium 231 </div> <div style="width: 15%;"> 92 U uranium 238 </div> <div style="width: 15%;"> 93 Np neptunium — </div> <div style="width: 15%;"> 94 Pu plutonium — </div> <div style="width: 15%;"> 95 Am americium — </div> <div style="width: 15%;"> 96 Cm curium — </div> <div style="width: 15%;"> 97 Bk berkelium — </div> <div style="width: 15%;"> 98 Cf californium — </div> <div style="width: 15%;"> 99 Es einsteinium — </div> <div style="width: 15%;"> 100 Fm fermium — </div> <div style="width: 15%;"> 101 Md mendelevium — </div> <div style="width: 15%;"> 102 No nobelium — </div> <div style="width: 15%;"> 103 Lr lawrencium — </div> </div>																																		

The volume of one mole of any gas is 24dm³ at room temperature and pressure (r.t.p.)

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