



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/22

Paper 2 Multiple Choice (Extended)

May/June 2016

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 1 3 0 1 7 6 6 3 5 3 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.
Do not use staples, paper clips, glue or correction fluid.
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.
DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.
Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
Any rough working should be done in this booklet.
A copy of the Periodic Table is printed on page 16.
Electronic calculators may be used.

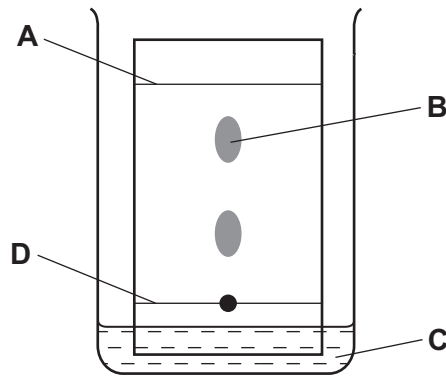
The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **16** printed pages.

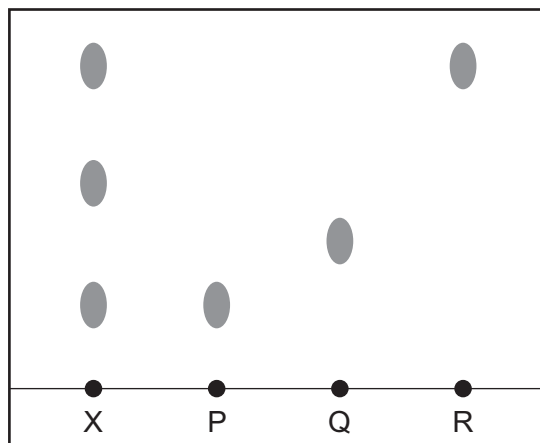
- 1 The particles of a substance gain energy and change from a regular ordered structure to a disordered structure with large distances between the particles.

Which change of state is described?

- A boiling
 B evaporation
 C melting
 D sublimation
- 2 In the chromatography experiment shown, which label represents the solvent front?



- 3 X is a mixture of colourless compounds. The diagram shows a chromatogram of X and of three pure compounds, P, Q and R.



Which statement is **not** correct?

- A A locating agent was used to develop the chromatogram of X.
 B P and R could be present in X.
 C P and R have different solubilities in the solvent.
 D Q has a greater R_f value than R.

4 Which statements about isotopes of the same element are correct?

- 1 They are atoms which have the same chemical properties because they have the same number of electrons in their outer shell.
- 2 They are atoms which have the same number of electrons and neutrons but different numbers of protons.
- 3 They are atoms which have the same number of electrons and protons but different numbers of neutrons.

A 1 and 2 **B** 1 and 3 **C** 2 only **D** 3 only

5 The table shows the electronic structure of four atoms.

atom	electronic structure
W	2,8,1
X	2,8,4
Y	2,8,7
Z	2,8,8

Which two atoms combine to form a covalent compound?

A W and X **B** W and Y **C** X and Y **D** X and Z

6 Which statement describes the attractive forces between molecules (intermolecular forces)?

- A** They are strong covalent bonds which hold molecules together.
- B** They are strong ionic bonds which hold molecules together.
- C** They are weak forces formed between covalently-bonded molecules.
- D** They are weak forces which hold ions together in a lattice.

7 Metals consist of a lattice of positive ions in a 'sea of electrons'.

Why is aluminium malleable?

- A** Its ions are attracted to the 'sea of electrons'.
- B** Its ions are tightly packed together.
- C** Its ions repel each other.
- D** Its layers of ions can slide over each other.

- 8 A sample of 16.0 g of a metal oxide, MO, is reduced to 12.8 g of the metal, M.

What is the relative atomic mass, A_r , of M?

- A 32 B 64 C 80 D 128

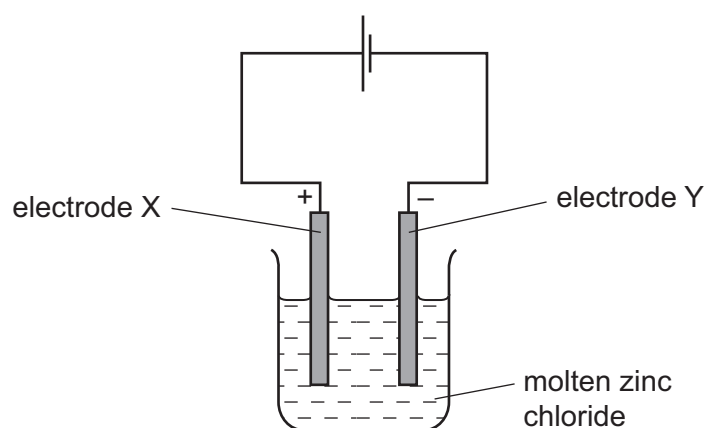
- 9 The equation for the reaction between calcium carbonate and hydrochloric acid is shown.



How many moles of calcium carbonate will give 24 cm^3 of carbon dioxide when reacted with an excess of the acid?

- A 1 mol B 0.1 mol C 0.01 mol D 0.001 mol

- 10 The diagram shows the electrolysis of molten zinc chloride, ZnCl_2 .

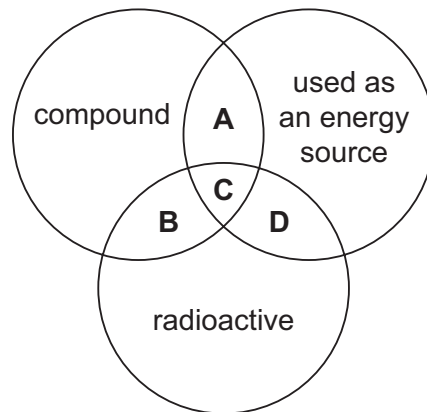


Which statement is correct?

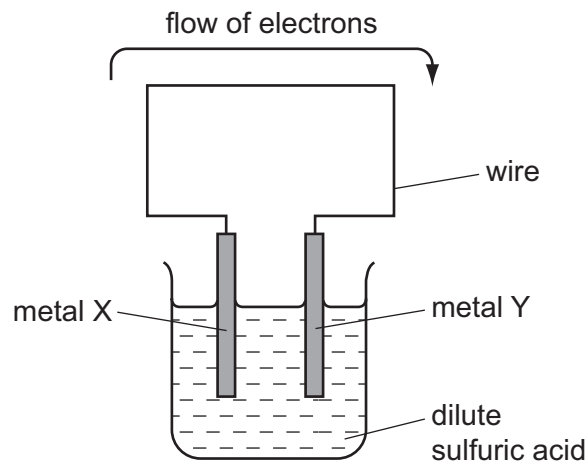
- A Oxidation occurs at electrode X and the equation is: $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$.
 B Oxidation occurs at electrode Y and the equation is: $\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$.
 C Reduction occurs at electrode X and the equation is: $\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$.
 D Reduction occurs at electrode Y and the equation is: $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$.

11 The diagram shows some properties that substances may have.

To which labelled part of the diagram does ^{235}U belong?



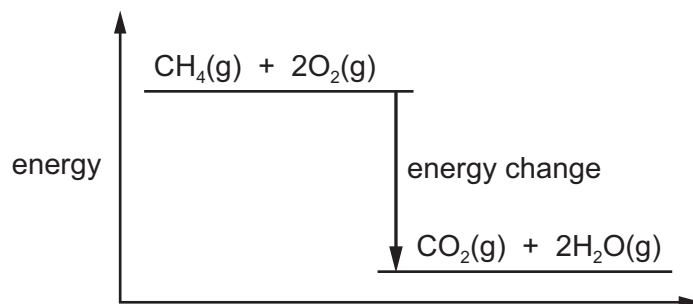
12 The diagram shows a simple cell.



For which pair of metals would electrons flow from metal X to metal Y?

	X	Y
A	copper	iron
B	copper	zinc
C	iron	zinc
D	zinc	iron

13 The energy level diagram for the combustion of methane is shown.

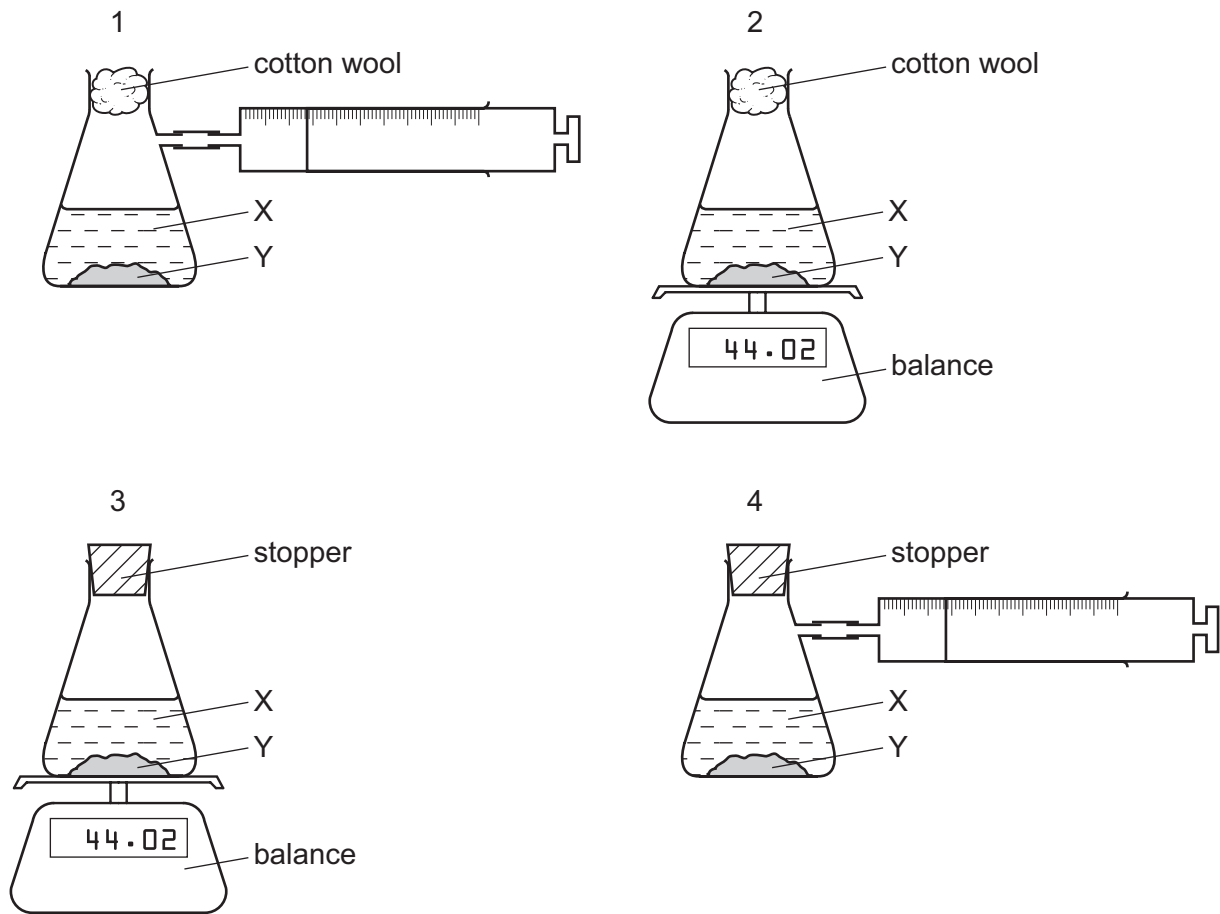


Which row gives the equation and energy change for this reaction?

	equation	energy change in kJ/mol
A	$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$	+891
B	$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$	-891
C	$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$	+891
D	$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$	-891

14 A liquid X reacts with solid Y to form a gas.

Which two diagrams show suitable methods for investigating the rate (speed) of the reaction?



A 1 and 3

B 1 and 4

C 2 and 3

D 2 and 4

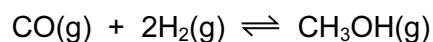
15 Which row describes how the energy of collision between particles changes when concentration and temperature are increased?

	concentration	temperature
A	increases	increases
B	increases	no change
C	no change	increases
D	no change	no change

16 Methanol is made by reacting carbon monoxide with hydrogen.

The reaction is exothermic and is a chemical equilibrium.

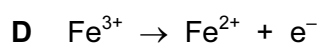
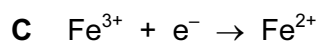
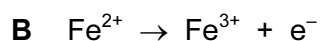
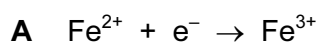
The equation for the reaction is shown.



Which changes in temperature and pressure increase the yield of methanol?

	temperature	pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

17 Which equation represents a reduction reaction?



18 Which statements are properties of an acid?

1 reacts with ammonium sulfate to form ammonia

2 turns red litmus blue

	1	2
A	✓	✓
B	✓	x
C	x	✓
D	x	x

19 Which row describes whether an amphoteric oxide reacts with acids and bases?

	reacts with acids	reacts with bases
A	no	no
B	no	yes
C	yes	no
D	yes	yes

20 Silver chloride is insoluble in water and is prepared by precipitation.

Which two substances can be used to make silver chloride?

- A** barium chloride and silver nitrate
- B** hydrochloric acid and silver
- C** hydrochloric acid and silver bromide
- D** sodium chloride and silver iodide

21 Where in the Periodic Table is the metallic character of the elements greatest?

	left or right side of a period	at the top or bottom of a group
A	left	bottom
B	left	top
C	right	bottom
D	right	top

22 Rubidium is a Group I metal.

Which statement about rubidium is **not** correct?

- A** It has a higher melting point than lithium.
- B** It has one electron in its outer shell.
- C** It reacts vigorously with water.
- D** It reacts with chlorine to form rubidium chloride, RbCl .

23 The table gives information about four elements, P, Q, R and S.

	melting point in °C	electrical conductivity of element when solid	density in g/cm ³	colour of iodide of element
P	98	good	0.97	white
Q	-39	good	13.53	red
R	1410	poor	2.33	colourless
S	1535	good	7.87	green

Which elements could be transition elements?

- A P, Q and S B Q and S only C R and S only D S only

24 Part of the Periodic Table is shown.

Which element is a gas that does **not** form a compound with potassium?

25 Some magnesium compounds undergo thermal decomposition.

What are the products of thermal decomposition of magnesium nitrate, Mg(NO₃)₂, and magnesium hydroxide, Mg(OH)₂?

	Mg(NO ₃) ₂	Mg(OH) ₂
A	MgO, NO ₂ and O ₂	MgO and H ₂ O
B	MgO, NO ₂ and O ₂	MgO and H ₂
C	Mg(NO ₂) ₂ and O ₂	MgO and H ₂ O
D	Mg(NO ₂) ₂ and O ₂	MgO and H ₂

26 Which property is **not** considered a typical metallic property?

- A good conductor of heat
 B low melting point
 C malleable (can be hammered into shape)
 D strong

27 Iron from a blast furnace is treated with oxygen and with calcium oxide to make steel.

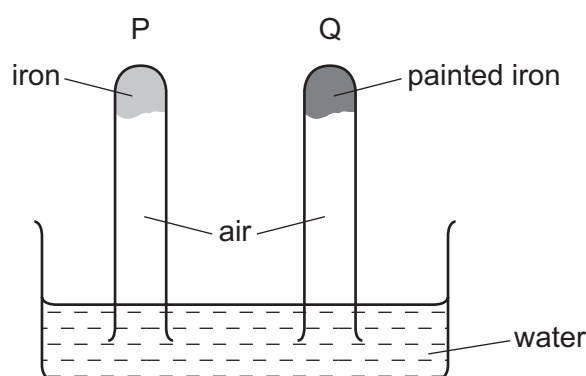
Which substances in the iron are removed?

	oxygen removes	calcium oxide removes
A	carbon	acidic oxides
B	carbon	basic oxides
C	iron	acidic oxides
D	iron	basic oxides

28 Why is cryolite used during the extraction of aluminium by electrolysis?

- A** It is a catalyst for the reaction.
- B** It lowers the melting point of the electrolyte.
- C** It protects the anodes.
- D** It separates the aluminium from the electrolyte.

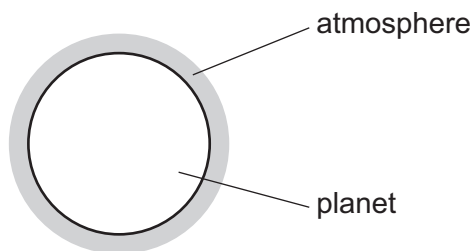
29 The diagram shows an experiment to investigate how paint affects the rusting of iron.



What happens to the water level in tubes P and Q?

	tube P	tube Q
A	falls	rises
B	no change	rises
C	rises	falls
D	rises	no change

30 A new planet has been discovered and its atmosphere has been analysed.



The table shows the composition of its atmosphere.

gas	percentage by volume
carbon dioxide	4
nitrogen	72
oxygen	24

Which gases are present in the atmosphere of the planet in a higher percentage than they are in the Earth's atmosphere?

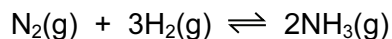
- A carbon dioxide and oxygen
- B carbon dioxide only
- C nitrogen and oxygen
- D nitrogen only

31 The gases coming from a car's engine contain oxides of nitrogen.

How are these oxides formed?

- A Nitrogen reacts with carbon dioxide.
- B Nitrogen reacts with carbon monoxide.
- C Nitrogen reacts with oxygen.
- D Nitrogen reacts with petrol.

32 Ammonia is manufactured by a reversible reaction.



The forward reaction is exothermic.

What is the effect of increasing the pressure on the percentage yield and rate of formation of ammonia?

	percentage yield	rate of formation
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

33 The Contact process is used for the manufacture of sulfuric acid.

Which statement about this process is **not** correct?

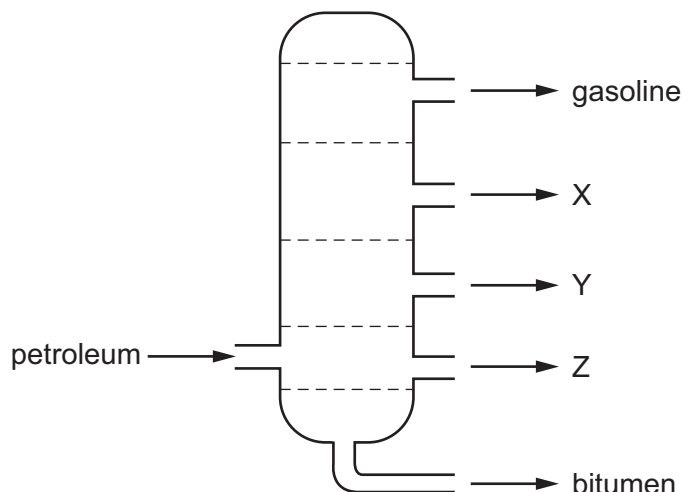
- A** A catalyst of iron is used.
- B** Oxygen from the air is used to react with sulfur dioxide.
- C** Sulfur trioxide dissolves in sulfuric acid to form oleum.
- D** The temperature used is around 450 °C.

34 Lime (calcium oxide) is used to treat waste water from a factory.

Which substance is removed by the lime?

- A** ammonia
- B** sodium chloride
- C** sodium hydroxide
- D** sulfuric acid

35 The diagram shows the separation of petroleum into fractions.



What could X, Y and Z represent?

	X	Y	Z
A	diesel oil	lubricating fraction	paraffin
B	lubricating fraction	diesel oil	paraffin
C	paraffin	lubricating fraction	diesel oil
D	paraffin	diesel oil	lubricating fraction

36 Which compound is **not** an alkane, C_nH_{2n+2} ?

- A** $CH_3CH_2CH_2CH_3$
- B** $(CH_3)_2CHCH_3$
- C** $CH_3CHCHCH_3$
- D** $(CH_3)_3CH$

37 An ester is formed when a carboxylic acid reacts with an alcohol.

Which ester is formed when propanoic acid and ethanol react?

- A** $CH_3CO_2CH_2CH_3$
- B** $CH_3CO_2CH_2CH_2CH_3$
- C** $CH_3CH_2CO_2CH_3$
- D** $CH_3CH_2CO_2CH_2CH_3$

38 What is an advantage of producing ethanol by fermentation of sugar compared to the catalytic addition of steam to ethene?

- A The alcohol produced is purer.
- B The process is faster.
- C The process uses high temperature.
- D The process uses renewable raw materials.

39 In which row are the monomer and polymer chain correctly matched?

	monomer	part of the polymer chain
A	$\text{CH}_3\text{CH}=\text{CHCH}_3$	$-\text{CH}(\text{CH}_3)-\text{CH}(\text{CH}_3)-\text{CH}(\text{CH}_3)-\text{CH}(\text{CH}_3)-$
B	$\text{CH}_2=\text{CHCl}$	$-\text{CHCl}-\text{CHCl}-\text{CHCl}-\text{CHCl}-$
C	$\text{CH}_3\text{CH}=\text{CH}_2$	$-\text{CH}_3-\text{CH}-\text{CH}_2-\text{CH}_3-\text{CH}-\text{CH}_2-$
D	$\text{CH}_2=\text{CHCH}_2\text{CH}_3$	$-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}(\text{CH}_2\text{CH}_3)-$

40 Which two polymers have the same linkages bonding the monomers together?

- A nylon and complex carbohydrate
- B nylon and protein
- C *Terylene* and complex carbohydrate
- D *Terylene* and protein

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The Periodic Table of Elements

		Group								
I	II	III	IV	V	VI	VII	VIII			
3 Li lithium 7	4 Be beryllium 9	<div style="border: 1px solid black; padding: 5px; width: fit-content; margin: auto;"> Key atomic number atomic symbol name relative atomic mass </div>								2 He helium 4
11 Na sodium 23	12 Mg magnesium 24									5 B boron 11
19 K potassium 39	20 Ca calcium 40	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40			
37 Rb rubidium 85	38 Sr strontium 88	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84			
55 Cs caesium 133	56 Ba barium 137	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131			
87 Fr francium —	88 Ra radium —	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —			
		29 Cu copper 64	28 Ni nickel 59	27 Co cobalt 59	26 Fe iron 56	25 Mn manganese 55	24 Cr chromium 52	23 V vanadium 51	22 Ti titanium 48	
		47 Ag silver 108	46 Pd palladium 106	45 Rh rhodium 103	44 Ru ruthenium 101	43 Tc technetium —	42 Mo molybdenum 96	41 Nb niobium 93	40 Zr zirconium 91	
		80 Hg mercury 201	79 Au gold 197	78 Pt platinum 195	77 Ir iridium 192	76 Os osmium 190	75 Re rhenium 186	74 W tungsten 184	73 Ta tantalum 181	
		112 Cn copernicium —	111 Rg roentgenium —	110 Ds darmstadtium —	109 Mt meitnerium —	108 Hs hassium —	107 Bh bohrium —	106 Sg seaborgium —	105 Db dubnium —	
		114 Fl flerovium —	113 Nh nihonium —	112 Cn copernicium —	111 Rg roentgenium —	110 Ds darmstadtium —	109 Mt meitnerium —	108 Hs hassium —	107 Bh bohrium —	
		68 Er erbium 167	67 Ho holmium 165	66 Dy dysprosium 163	65 Tb terbium 159	64 Gd gadolinium 157	63 Eu europium 152	62 Sm samarium 150	61 Pm promethium —	
		100 Fm fermium —	99 Es einsteinium —	98 Cf californium —	97 Bk berkelium —	96 Cm curium —	95 Am americium —	94 Pu plutonium —	93 Np neptunium —	
		71 Lu lutetium 175	70 Yb ytterbium 173	69 Tm thulium 169	68 Er erbium 167	67 Ho holmium 165	66 Dy dysprosium 163	65 Tb terbium 159	64 Gd gadolinium 157	
		103 Lr lawrencium —	102 No nobelium —	101 Md mendelevium —	100 Fm fermium —	99 Es einsteinium —	98 Cf californium —	97 Bk berkelium —	96 Cm curium —	
		57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	
		89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	

		68 Er erbium 167	67 Ho holmium 165	66 Dy dysprosium 163	65 Tb terbium 159	64 Gd gadolinium 157	63 Eu europium 152	62 Sm samarium 150	61 Pm promethium —	60 Nd neodymium 144	59 Pr praseodymium 141	58 Ce cerium 140	57 La lanthanum 139			
lanthanoids		71 Lu lutetium 175	70 Yb ytterbium 173	69 Tm thulium 169	68 Er erbium 167	67 Ho holmium 165	66 Dy dysprosium 163	65 Tb terbium 159	64 Gd gadolinium 157	63 Eu europium 152	62 Sm samarium 150	61 Pm promethium —	60 Nd neodymium 144	59 Pr praseodymium 141	58 Ce cerium 140	57 La lanthanum 139
actinoids		103 Lr lawrencium —	102 No nobelium —	101 Md mendelevium —	100 Fm fermium —	99 Es einsteinium —	98 Cf californium —	97 Bk berkelium —	96 Cm curium —	95 Am americium —	94 Pu plutonium —	93 Np neptunium —	92 U uranium 238	91 Pa protactinium 231	90 Th thorium 232	89 Ac actinium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.)