

Centre Number	Candidate Number	Name
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UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/03

Paper 3

May/June 2006

1 hour 15 minutes

Candidates answer on the Question Paper.
No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.
Write in dark blue or black pen.
You may use a pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.

Answer **all** questions.
A copy of the Periodic Table is printed on page 16.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
Total	

This document consists of **16** printed pages.

1 Iron is a transition element.

(a) Which of the following statements about transition elements are correct?

Tick **three** boxes.

The metals are highly coloured e.g. yellow, green, blue.

The metals have low melting points.

Their compounds are highly coloured.

Their compounds are colourless.

The elements and their compounds are often used as catalysts.

They have more than one oxidation state.

[3]

(b) (i) In which Period in the Periodic Table is iron to be found?

..... [1]

(ii) Use the Periodic Table to work out the number of protons and the number of neutrons in one atom of iron.

number of protons = number of neutrons = [1]

(c) Iron is extracted in a blast furnace. The list below gives some of the substances used or formed in the extraction.

carbon monoxide coke iron ore limestone slag

(i) Which substance is a mineral containing largely calcium carbonate?

..... [1]

(ii) Which substance is formed when impurities in the ore react with calcium oxide?

..... [1]

(iii) Which substance is also called hematite?

..... [1]

(d) State **two** functions of the coke used in the blast furnace.

.....
..... [2]

(e) Most of the iron is converted into mild steel or stainless steel. Give **one** use for each.

mild steel
stainless steel [2]

2 Some reactions of metals **W**, **X**, **Y** and **Z** are given below.

metal	reaction with water	reaction with dilute hydrochloric acid
W	A few bubbles form slowly in cold water.	Vigorous reaction. Gas given off.
X	Vigorous reaction. Metal melts. Gas given off.	Explosive reaction. Should not be attempted.
Y	No reaction.	No reaction.
Z	Does not react with cold water. Hot metal reacts with steam.	Steady fizzing.

(a) Arrange these metals in order of reactivity.

most reactive

.....

.....

least reactive [2]

(b) Which of these metals could be

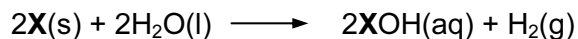
(i) magnesium,

..... [1]

(ii) copper?

..... [1]

(c) The equation for the reaction of **X** with cold water is given below.



(i) Describe the test you would use to show that the gas evolved is hydrogen.

.....
..... [1]

(ii) How could you show that the water contained a compound of the type **XOH**?

.....
..... [2]

(iii) In which group of the Periodic Table does metal **X** belong?

..... [1]

(iv) The ore of **X** is its chloride. Suggest how metal **X** could be extracted from its chloride.

..... [2]

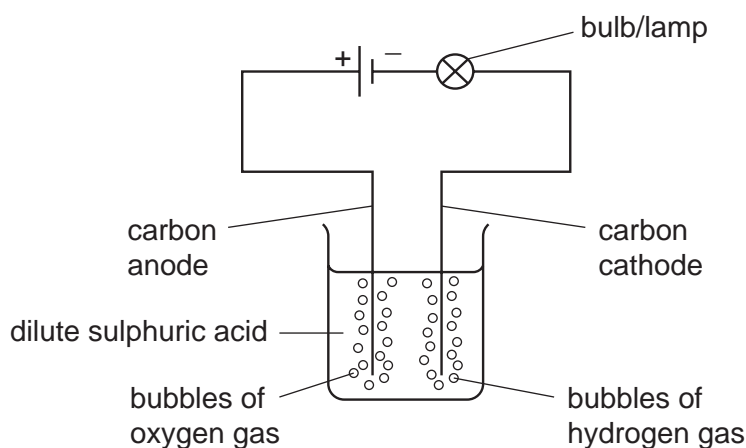
- 3 (a) Four bottles were known to contain aqueous ammonia, dilute hydrochloric acid, sodium hydroxide solution and vinegar, which is dilute ethanoic acid. The bottles had lost their labels. The pH values of the four solutions were 1, 4, 10 and 13.

Complete the table.

solution	pH
aqueous ammonia	
dilute hydrochloric acid	
sodium hydroxide solution	
vinegar	

[2]

- (b) The following apparatus was set up to investigate the electrical conductivity of dilute acids.

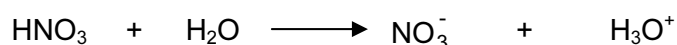


Dilute sulphuric acid is a strong acid. If it was replaced by a weak acid, what **two** differences in the observations would you expect to make?

.....

 [2]

- (c) When nitric acid is added to water the following reaction occurs.



Give the name and the formula of the particle which is transferred from nitric acid to water.

name

formula

[2]

(d) This question is concerned with the following oxides.

aluminium oxide Al_2O_3

calcium oxide CaO

carbon dioxide CO_2

carbon monoxide CO

magnesium oxide MgO

sulphur dioxide SO_2

(i) Which of the above oxides will react with hydrochloric acid but not with aqueous sodium hydroxide?

..... [1]

(ii) Which of the above oxides will react with aqueous sodium hydroxide but not with hydrochloric acid?

..... [1]

(iii) Which of the above oxides will react both with hydrochloric acid and with aqueous sodium hydroxide?

..... [1]

(iv) Which of the above oxides will react neither with hydrochloric acid nor with aqueous sodium hydroxide?

..... [1]

4 The first three elements in Group IV are
carbon,
silicon,
germanium.

(a) The element germanium has a diamond-type structure. Describe the structure of germanium. A diagram is acceptable.

[2]

(b) Unlike diamond, graphite is soft and is a good conductor of electricity.

(i) Explain why graphite has these properties.

.....

 [3]

(ii) Give a use of graphite that depends on one of these properties.

property

use [1]

(c) Carbon dioxide and silicon(IV) oxide have similar formulae but different types of structure.

(i) Give the formulae of these oxides.

..... [1]

(ii) How are their structures different?

.....
 [2]

(d) All these elements form compounds with hydrogen called hydrides. The saturated hydrides of carbon are the alkanes. Predict the formula of the hydride of germanium which contains two germanium atoms.

..... [1]

5 Sulphuric acid is made by the Contact process in the following sequence of reactions.



(a) (i) How is sulphur dioxide made from sulphur?

..... [1]

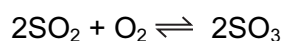
(ii) Sulphur dioxide has other uses.
Why is it used in the manufacture of paper?

..... [1]

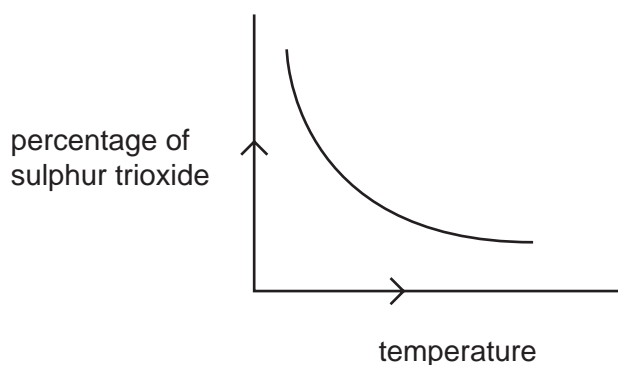
(iii) How does it preserve food?

..... [1]

(b) The equation for a stage of the Contact process is



The percentage of sulphur trioxide in the equilibrium mixture varies with temperature.



(i) How does the percentage of sulphur trioxide in the equilibrium mixture vary as the temperature increases? Circle the correct answer.

increases

stays the same

decreases

[1]

(ii) Is the forward reaction in the equilibrium $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$ exothermic or endothermic? Give a reason for your choice.

.....
..... [2]

(iii) Explain, mentioning both rate and percentage yield, why the temperature used in the Contact process is 450°C.

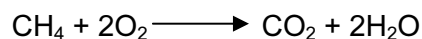
.....
..... [2]

(iv) Describe how the sulphur trioxide is changed into concentrated sulphuric acid.

.....
..... [2]

- 6 (a) Exothermic reactions produce heat energy.

An important fuel is methane, natural gas. The equation for its combustion is as follows.



- (i) In chemical reactions bonds are broken and new bonds are formed. Using this reaction give an example of

a bond that is broken,

a bond that is formed. [2]

- (ii) Explain, using the idea of bonds forming and breaking, why this reaction is exothermic, that is it produces heat energy.

.....
..... [2]

- (b) Some radioactive isotopes are used as nuclear fuels.

- (i) Give the symbol and the nucleon number of an isotope that is used as a nuclear fuel.

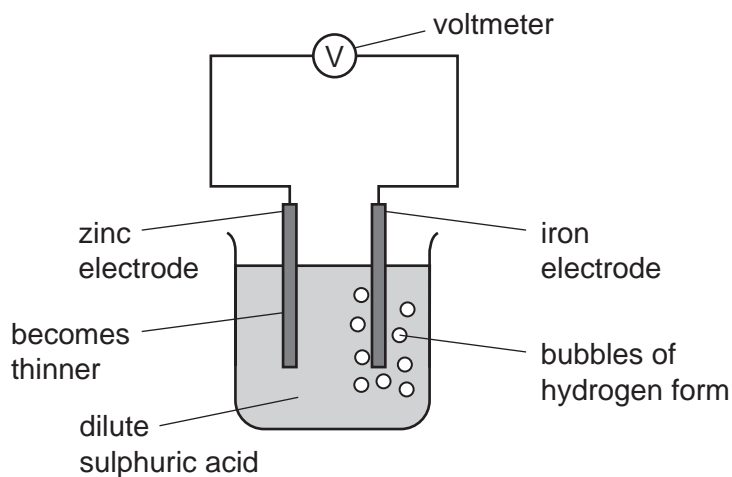
..... [2]

- (ii) Give another use of radioactive isotopes.

..... [1]

(c) Cell reactions are both exothermic and redox. They produce electrical energy as well as heat energy.

(i) The diagram shows a simple cell.



Which substance in this cell is the reductant and which ion is the oxidant?

reductant

oxidant [2]

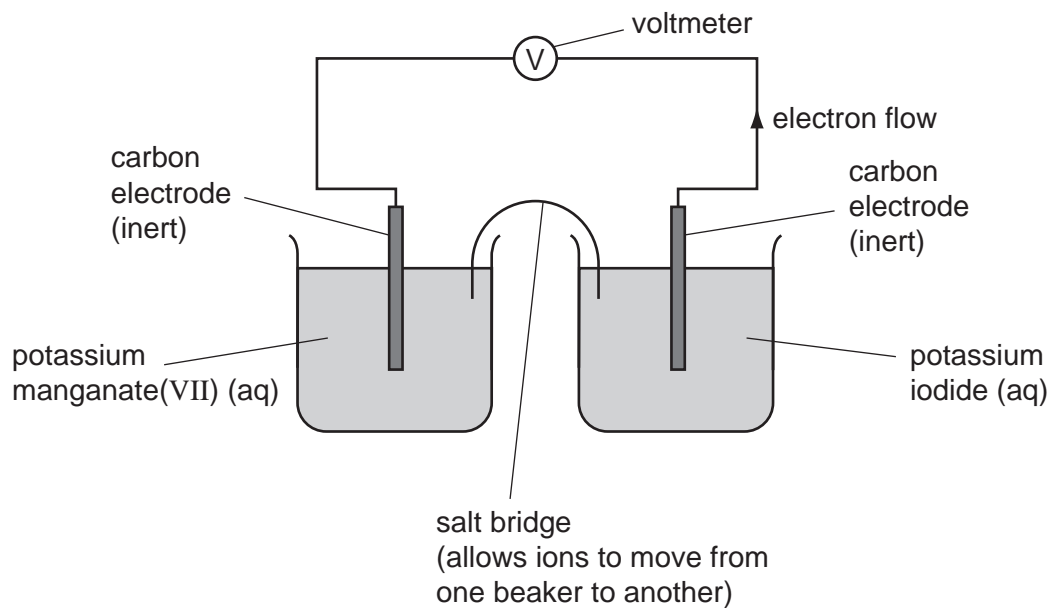
(ii) How could the voltage of this cell be increased?

..... [1]

(iii) What is the important large scale use, relating to iron and steel, of this type of cell reaction?

..... [1]

(d) Cells can be set up with inert electrodes and the electrolytes as oxidant and reductant.



The potassium manganate(VII) is the oxidant and the potassium iodide is the reductant.

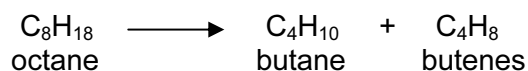
(i) Describe the colour change that would be observed in the left hand beaker.

..... [2]

(ii) Write an ionic equation for the reaction in the right hand beaker.

..... [2]

- 7 The fractional distillation of crude oil usually produces large quantities of the heavier fractions. The market demand is for the lighter fractions and for the more reactive alkenes. The heavier fractions are cracked to form smaller alkanes and alkenes as in the following example.



- (a) (i) Write a different equation for the cracking of octane.



- (ii) The cracking of octane can produce isomers with the molecular formula C_4H_8 . Draw the structural formulae of two of these isomers.

[2]

- (b) (i) Give the essential condition for the reaction between chlorine and butane.

..... [1]

- (ii) What type of reaction is this?

..... [1]

- (iii) This reaction produces a mixture of products. Give the names of **two** products that contain four carbon atoms per molecule.

..... and [2]

- (c) Alkenes are more reactive than alkanes and are used to make a range of organic chemicals. Propene, $\text{CH}_3\text{-CH=CH}_2$, is made by cracking. Give the structural formula of the addition product when propene reacts with the following.

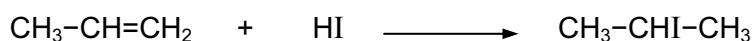
(i) water

[1]

(ii) bromine

[1]

- (d) Propene reacts with hydrogen iodide to form 2-iodopropane.



1.4 g of propene produced 4.0 g of 2-iodopropane.

Calculate the percentage yield.

moles of $\text{CH}_3\text{-CH=CH}_2$ reacted =

maximum moles of $\text{CH}_3\text{-CHI-CH}_3$ that could be formed =

mass of one mole of $\text{CH}_3\text{-CHI-CH}_3 = 170 \text{ g}$

maximum mass of 2-iodopropane that could be formed =

percentage yield%

[4]

DATA SHEET
The Periodic Table of the Elements

		Group										0																							
		I	II	III	IV	V	VI	VII																											
		<table border="1" style="width: 100%; border-collapse: collapse; text-align: center;"> <tr> <td>1</td> <td>H Hydrogen 1</td> </tr> </table>										1	H Hydrogen 1	4	He Helium 2																				
1	H Hydrogen 1																																		
7	Li Lithium 3	9	Be Beryllium 4											20	Ne Neon 10																				
23	Na Sodium 11	24	Mg Magnesium 12											32	Ar Argon 18																				
39	K Potassium 19	40	Ca Calcium 20	45	Sc Scandium 21	48	Ti Titanium 22	51	V Vanadium 23	52	Cr Chromium 24	55	Mn Manganese 25	56	Fe Iron 26	59	Co Cobalt 27	59	Ni Nickel 28	64	Cu Copper 29	65	Zn Zinc 30	70	Ga Gallium 31	73	Ge Germanium 32	75	As Arsenic 33	79	Se Selenium 34	80	Br Bromine 35	84	Kr Krypton 36
85	Rb Rubidium 37	88	Sr Strontium 38	89	Y Yttrium 39	91	Zr Zirconium 40	93	Nb Niobium 41	96	Mo Molybdenum 42	101	Ru Ruthenium 44	103	Rh Rhodium 45	106	Pd Palladium 46	108	Ag Silver 47	112	Cd Cadmium 48	115	In Indium 49	119	Sn Tin 50	122	Sb Antimony 51	127	I Iodine 53	131	Xe Xenon 54				
133	Cs Caesium 55	137	Ba Barium 56	139	La Lanthanum 57	178	Hf Hafnium 72	181	Ta Tantalum 73	184	W Tungsten 74	190	Os Osmium 76	192	Ir Iridium 77	195	Pt Platinum 78	197	Au Gold 79	201	Hg Mercury 80	204	Tl Thallium 81	207	Pb Lead 82	209	Bi Bismuth 83	210	Po Polonium 84	210	Rn Radon 86				
226	Fr Francium 87	226	Ra Radium 88	227	Ac Actinium 89											227	Fr Francium 87																		

140	Ce Cerium 58	141	Pr Praseodymium 59	144	Nd Neodymium 60	146	Pm Promethium 61	150	Sm Samarium 62	152	Eu Europium 63	157	Gd Gadolinium 64	162	Dy Dysprosium 66	165	Ho Holmium 67	167	Er Erbium 68	169	Tm Thulium 69	173	Yb Ytterbium 70	175	Lu Lutetium 71
232	Th Thorium 90	232	Pa Protactinium 91	238	U Uranium 92	238	Np Neptunium 93	238	Pu Plutonium 94	238	Am Americium 95	238	Cm Curium 96	238	Cf Californium 98	238	Es Einsteinium 99	238	Fm Fermium 100	238	Md Mendelevium 101	238	No Nobelium 102	238	Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X
b	b

a = relative atomic mass
 X = atomic symbol
 b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).