

Please check the examination details below before entering your candidate information

Candidate surname

Other names

Centre Number

Candidate Number

--	--	--	--	--

--	--	--	--

Pearson Edexcel International Advanced Level

Time 1 hour 20 minutes

Paper
reference

WCH16/01

Chemistry

International Advanced Level

UNIT 6: Practical Skills in Chemistry II

You must have:

Scientific calculator

Total Marks

Instructions

- Use **black** ink or **black** ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 50.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- You will be assessed on your ability to organise and present information, ideas, descriptions and arguments clearly and logically, including your use of grammar, punctuation and spelling.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Check your answers if you have time at the end.

Turn over ►

P67132A

©2021 Pearson Education Ltd.

E:1/1/1/1/1/



Pearson

Answer ALL questions. Write your answers in the spaces provided.

1 This question is about copper and some of its compounds.

(a) Two tests were carried out on separate samples of an aqueous solution of copper(II) sulfate.

(i) **Test 1**

A few drops of aqueous sodium hydroxide were added to a sample of the copper(II) sulfate solution.

State what you would see.

(1)

(ii) **Test 2**

A few drops of concentrated hydrochloric acid were added to another sample of the copper(II) sulfate solution.

More of the concentrated hydrochloric acid was added until it was present in excess.

Describe the changes that would be observed during this test.

(2)

(b) Describe a test, and its positive result, to confirm the presence of the sulfate ion in another sample of the copper(II) sulfate solution.

(2)

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



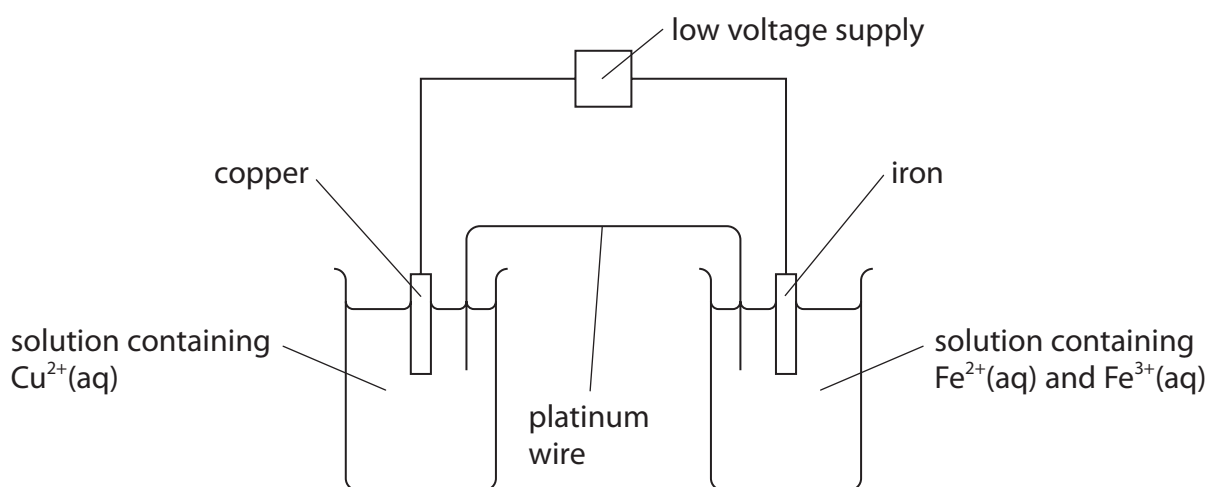
- (c) An electrochemical cell was made from the electrode systems represented by these half-equations:



- (i) Calculate $E_{\text{cell}}^{\ominus}$ for the electrochemical cell.

(1)

- (ii) A student drew a diagram of an experiment to measure the standard emf of the cell.



Identify three mistakes in this diagram and the changes needed to correct them.

Assume that standard conditions were used.

(3)

Mistake	Change needed to correct mistake

- (d) Brass is an alloy of copper and zinc.
A student determined the percentage of copper in a sample of brass.

Procedure

- weigh the sample of brass
 - place the brass in a beaker and add concentrated nitric acid until all the brass dissolves
 - transfer the solution and washings to a 250.0 cm^3 volumetric flask
 - make the solution up to the mark with distilled water and mix well
 - pipette 25.0 cm^3 of the solution into a conical flask
 - neutralise the excess nitric acid in the solution
 - add 10 cm^3 of potassium iodide solution (an excess) to the conical flask
 - titrate the iodine produced with $0.100 \text{ mol dm}^{-3}$ sodium thiosulfate solution using starch indicator
 - repeat the titration until concordant titres are obtained.
- (i) Copper and zinc both react with concentrated nitric acid to form the metal nitrates, nitrogen dioxide and water.

Write the balanced equation for the reaction of zinc with concentrated nitric acid.

State symbols are not required.

(1)

- (ii) Name the most suitable piece of apparatus to measure the 10 cm^3 of potassium iodide solution.

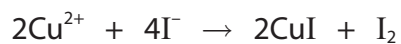
(1)

- (iii) State at what point in the titration the starch solution should be added.

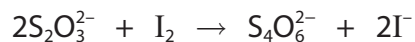
(1)



(iv) Only Cu^{2+} ions in the solution react with the aqueous potassium iodide.



The iodine reacts with sodium thiosulfate solution.



Results

Mass of brass sample = 3.90 g

Mean titre of $0.100 \text{ mol dm}^{-3}$ sodium thiosulfate solution = 28.60 cm^3

Calculate the percentage, by mass, of copper in this sample of brass.
Give your answer to an appropriate number of significant figures.

(5)

(Total for Question 1 = 17 marks)



2 Two organic compounds, **A** and **B**, are colourless liquids.

Each compound contains only **one** functional group.

(a) Two tests were carried out on **A**. The observation for each test was recorded in the table.

(i) Complete the statements in the inference column by writing the names or formulae of the functional groups.

(2)

Test	Observation	Inference
Test 1 A few drops of A were added to 2 cm ³ of a solution of 2,4-dinitrophenylhydrazine (Brady's reagent)	An orange precipitate formed	A could contain or
Test 2 A few drops of A were added to 2 cm ³ of Fehling's solution The mixture was warmed in a water bath	A red precipitate formed	The functional group present in A is

(ii) Give the name or formula of the red precipitate formed in **Test 2**.

(1)

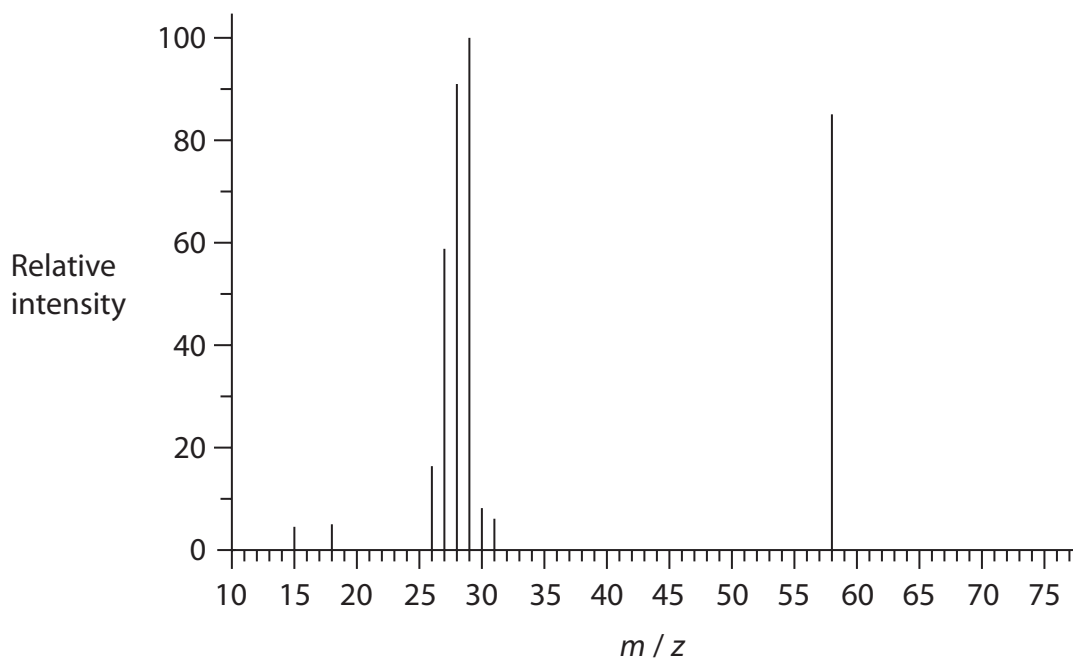


DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

(b) A simplified mass spectrum of **A** is shown.



(i) Give the formula of **one** of the ions responsible for the peak at $m/z = 29$.

(1)

(ii) **A** contains one functional group.

Give the m/z value of the molecular ion and the structure of **A**.

(1)

m/z value of the molecular ion

structure of **A**



P 6 7 1 3 2 A 0 7 1 6

(c) Two tests were carried out on **B**.

(i) Complete the statements in the observation and inference columns.

(2)

Test	Observation	Inference
<p>Test 3</p> <p>2 drops of B were dissolved in 2 cm³ of water</p> <p>A few drops of Universal Indicator were added to the solution</p>	<p>The colour of the mixture was</p>	<p>The solution is alkaline</p>
<p>Test 4</p> <p>B was added drop by drop to aqueous copper(II) sulfate until B was present in excess</p>	<p>A pale blue precipitate formed with the first few drops of B</p> <p>This dissolved to form a deep blue solution when excess B was added</p>	<p>The name of the functional group in B is</p>

(ii) **B** has a molar mass of 59 g mol⁻¹.

Suggest a structure for **B**.

(1)

(Total for Question 2 = 8 marks)



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



- 3 A student carried out an experiment to determine the enthalpy change when solid lithium chloride, LiCl, dissolved in water to form a solution.

Procedure

Step 1 Use a pipette to place 25.0 cm^3 of distilled water into a polystyrene cup.

Step 2 Measure and record the initial temperature of the water.

Step 3 Add 2.12 g of lithium chloride to the water.

Step 4 Stir the mixture and record the highest temperature reached.

- (a) Give a reason why a polystyrene cup was used instead of a glass beaker in Step 1. (1)

- (b) The temperature rise was $12.5 \text{ }^\circ\text{C}$.

Calculate the enthalpy change for the formation of this solution of lithium chloride.

Include a sign and units in your answer.

[Assume: specific heat capacity of the solution = $4.18 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$
density of the solution = 1.00 g cm^{-3}]

(3)



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

- (c) The thermometer used to measure the temperature change had an uncertainty of $\pm 0.25\text{ }^\circ\text{C}$ for each measurement.

Calculate the percentage uncertainty in the temperature **change** in this experiment.

(1)

- (d) The temperature rise in this experiment was lower than expected, due to heat loss to the surroundings.

Describe changes to the procedure that would give a more accurate temperature rise.

Include the use of a stopwatch and details of a graph you would plot.

(5)

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

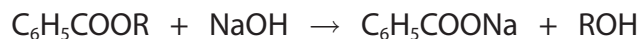
(Total for Question 3 = 10 marks)



4 This question is about the alkaline hydrolysis of an ester, **X**.

X is an alkyl benzoate and can be represented by the formula C_6H_5COOR , where R is the alkyl group.

The equation for the hydrolysis is



Procedure

Step 1 Measure 5.0 cm^3 of **X** and pour it into a pear-shaped flask. Add 25 cm^3 (an excess) of aqueous sodium hydroxide solution and a few anti-bumping granules.

Step 2 Heat the flask and contents under reflux for 20 minutes.

Step 3 Allow the apparatus to cool and then rearrange it for distillation. Distil the mixture and collect about 2 cm^3 of the alcohol ROH.

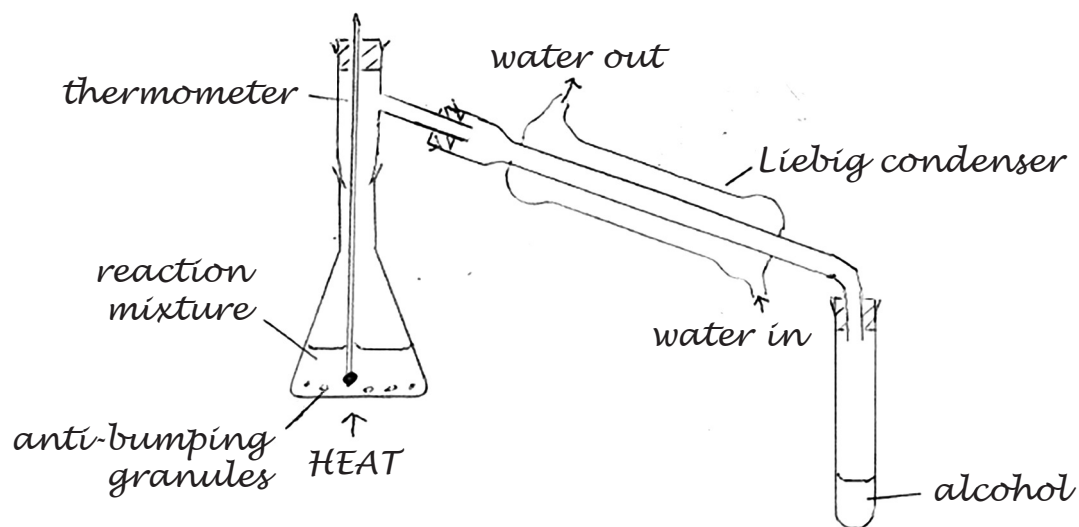
Step 4 Allow the pear-shaped flask to cool, pour the contents into a beaker and add excess dilute hydrochloric acid. Impure benzoic acid forms as crystals in the mixture.

Step 5 Recrystallise the benzoic acid using water as the solvent.

Step 6 Weigh the dry crystals and determine their melting temperature.



- (a) A student drew a diagram of the apparatus set up for distillation in Step 3. There are three errors in the diagram. Assume the apparatus is clamped correctly and an appropriate heat source is used.



Identify the three errors and how they should be corrected.

(3)

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....



(b) The distillate collected in Step 3 is the alcohol ROH.

Describe a **chemical** test, and its positive result, to show the presence of an -OH group in **any** alcohol.

(2)

(c) (i) Write an equation for the reaction taking place in Step 4.

Use structural formulae for the organic substances. State symbols are not required.

(1)

(ii) State what should be done to separate the benzoic acid from the mixture produced in Step 4, before carrying out Step 5.

(1)

(d) Describe the **first** stage in the recrystallisation process in Step 5.

(1)

(e) The melting temperature of pure benzoic acid is 122 °C.

State **two** ways in which the melting temperature changes if the benzoic acid is **not** pure.

(2)



(f) The molar mass of **X**, C_6H_5COOR , is 178 g mol^{-1} .

(i) Deduce the formula of the alkyl group, R.

(1)

(ii) Use your answer to (f)(i) to draw the structures of the four possible alcohols, ROH.

(2)

(iii) The part of the ^{13}C NMR spectrum of **X** corresponding to the R group contains only two peaks.

Deduce the structure of **X**.

(2)

(Total for Question 4 = 15 marks)

TOTAL FOR PAPER = 50 MARKS

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



The Periodic Table of Elements

1 2 3 4 5 6 7 0 (8)
(18)

1.0	H	hydrogen	1
-----	----------	----------	---

Key

relative atomic mass
atomic symbol
name
atomic (proton) number

(1)	(2)	(3)	(4)	(5)	(6)	(7)	(8)	(9)	(10)	(11)	(12)	(13)	(14)	(15)	(16)	(17)	(18)
6.9	9.0	45.0	47.9	50.9	52.0	54.9	55.8	58.9	58.7	63.5	65.4	10.8	12.0	14.0	16.0	19.0	4.0
Li	Be	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	B	C	N	O	F	He
lithium	beryllium	scandium	titanium	vanadium	chromium	manganese	iron	cobalt	nickel	copper	zinc	boron	carbon	nitrogen	oxygen	fluorine	helium
3	4	21	22	23	24	25	26	27	28	29	30	5	6	7	8	9	2
23.0	24.3	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	27.0	28.1	31.0	32.1	35.5	39.9
Na	Mg	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	Al	Si	P	S	Cl	Ar
sodium	magnesium	yttrium	zirconium	niobium	molybdenum	technetium	ruthenium	rhodium	palladium	silver	cadmium	aluminium	silicon	phosphorus	sulfur	chlorine	argon
11	12	39	40	41	42	43	44	45	46	47	48	13	14	15	16	17	18
39.1	40.1	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	69.7	72.6	74.9	79.0	79.9	83.8
K	Ca	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Ga	Ge	As	Se	Br	Kr
potassium	calcium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	gallium	germanium	arsenic	selenium	bromine	krypton
19	20	57	72	73	74	75	76	77	78	79	80	31	32	33	34	35	36
85.5	87.6	138.9	178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	69.7	72.6	74.9	79.0	79.9	83.8
Rb	Sr	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	In	Sn	Sb	Te	I	Xe
rubidium	strontium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	indium	tin	antimony	tellurium	iodine	xenon
37	38	57	72	73	74	75	76	77	78	79	80	49	50	51	52	53	54
132.9	137.3	138.9	178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	114.8	118.7	121.8	127.6	126.9	131.3
Cs	Ba	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Pb	Bi	Po	At	Rn	Rn
caesium	barium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	lead	bismuth	polonium	astatine	radon	radon
55	56	57	72	73	74	75	76	77	78	79	80	82	83	84	85	86	86
[223]	[226]	[227]	[261]	[262]	[266]	[264]	[277]	[268]	[271]	[272]	[272]	204.4	207.2	209.0	[210]	[222]	[222]
Fr	Ra	Ac*	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Rg	Tl	Pb	Bi	Po	At	Rn
francium	radium	actinium	rutherfordium	dubnium	seaborgium	bohrium	hassium	meitnerium	darmstadtium	roentgenium	roentgenium	thallium	lead	bismuth	polonium	astatine	radon
87	88	89	104	105	106	107	108	109	110	111	111	81	82	83	84	85	86

Elements with atomic numbers 112-116 have been reported but not fully authenticated

140	141	144	150	152	157	163	165	167	169	173	175
Ce	Pr	Nd	Sm	Eu	Gd	Dy	Ho	Er	Tm	Yb	Lu
cerium	praseodymium	neodymium	samarium	europium	gadolinium	dysprosium	holmium	erbium	thulium	ytterbium	lutetium
58	59	60	62	63	64	66	67	68	69	70	71
232	[231]	238	[242]	[243]	[247]	[251]	[254]	[253]	[256]	[254]	[257]
Th	Pa	U	Pu	Am	Cm	Cf	Es	Fm	Md	No	Lr
thorium	protactinium	uranium	plutonium	americium	curium	californium	einsteinium	fermium	mendeleevium	nobelium	lawrencium
90	91	92	94	95	96	98	99	100	101	102	103

* Lanthanide series

* Actinide series

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

