



Mark Scheme (Results)

October 2018

Pearson Edexcel International

Advanced Level

In Chemistry (WCH06)

Paper 01 Chemistry Laboratory Skills II

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

() means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer.

ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

Question Number	Answer	Reject	Mark
1 (a) (i)	Observation with copper(II) sulfate and with copper(II) hydroxide: (Dissolves to form a) blue solution Allow Blue-green solution Ignore just blue solid	Green / yellow solution	1

Question Number	Answer	Reject	Mark
1 (a) (ii)	Observation with copper(II) sulfate: White precipitate/solid/crystals and Observation starting with copper(II) hydroxide: No change/No precipitate/Remain blue (solution) ALLOW No observation/No reaction If both copper(II) sulfate observations are correct, then 1 mark out of two for parts (i) and (ii)	Solid dissolves Just 'No'	1

Question Number	Answer	Reject	Mark
1 (b) (i)	(Observation) Pink (solution) ALLOW Shades of pink (Inference) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ALLOW $[\text{Co}(\text{H}_2\text{O})_4]^{2+}$ IGNORE lack of square brackets Mark independently	(1) Purple $[\text{Co}(\text{H}_2\text{O})_6\text{Cl}_2]^{2+}$ (1) Just Co^{2+} $\text{Co}^{2+}(\text{aq})$	2

Question Number	Answer	Reject	Mark
1 (b)(ii)	<p>(To an aqueous solution) Add nitric acid/HNO₃ and silver nitrate/AgNO₃ (solution)</p> <p>If name and formula given, then both must be correct ALLOW reagents to be given in either order but not after the white ppt observation (1)</p> <p>White precipitate/solid</p> <p>IGNORE Addition of aqueous ammonia (1)</p> <p>ALLOW To the solids Add concentrated sulfuric acid/H₂SO₄ (1)</p> <p>Steamy/Misty fumes (1)</p> <p>OR</p> <p>ALLOW To the solids Add acidified manganate(VII) (1) Bubbles (of chlorine gas) (1)</p>	Use of HCl scores (0)	2

Question Number	Answer	Reject	Mark
1 (c)	<p>(Acidified / H⁺) MnO₄⁻ / (Acidified / H⁺ potassium) manganate((VII)) / permanganate ALLOW (H₂SO₄) KMnO₄ (1)</p> <p>$E^{\circ}_{\text{cell}} = (+) 0.51(\text{V})$</p> <p>No TE on incorrect reagent (1)</p>	<p>Additional reagents</p> <p>K₂MnO₄</p> <p>0.5 -0.51</p>	2

Question Number	Answer	Reject	Mark
1 (d)(i)	<p>VOSO₄.4H₂O/ VOSO₄.3H₂O/ VOSO₄.2H₂O/ VOSO₄.H₂O</p> <p>ALLOW VOSO₄.5H₂O</p>	VOSO ₄	1

Question Number	Answer	Reject	Mark
1 (d)(ii)	Not all of the water (of crystallisation) had been removed (from some of the sample) OR (More) water evaporates / was driven off ALLOW Steam given off	Loss of O ₂	1

Question Number	Answer	Reject	Mark
1 (d)(iii)	Molar mass = 253 (g mol ⁻¹) (1) % water of crystallisation = % loss in mass = 100 x 90/253 = 35.573/ 35.57/35.6/36 (%) Ignore SF except 1 TE on molar mass provided % loss is less than 100 (1) Correct final answer without working scores (2)		2

(Total for Question 1 = 12 marks)

Question Number	Acceptable Answers	Reject	Mark
2(a)	The solid needs to dissolve in the reaction mixture (before it can quench) ALLOW The reaction is quenched/stopped by the solution quicker Ignore references to just 'surface area greater' ALLOW Comparison of the reaction rate of the solid being slower to that of the solution		1

Question Number	Acceptable Answers	Reject	Mark
2(b)(i)	(Estimated volume) 20.9(0) to 21.1(0) (cm ³) Answer to 1dp ALLOW Range within that stated above, such as 20.9(0) – 21.0(0)	21	1

Question Number	Acceptable Answers	Reject	Mark														
2(b)(ii)	<p>Graph M1 Axes correct with sensible scales so at least half of the graph paper on both axes is covered (1)</p> <p>Graph M2 Axes labels fully correct with units ALLOW Volume/cm³ (1)</p> <p>Graph M3 All points plotted correctly (± 1 small square) and straight line (1)</p> <p>Example graph</p> <table border="1"> <caption>Data points from the example graph</caption> <thead> <tr> <th>Time / minutes</th> <th>Volume of sodium bisulfate / cm³</th> </tr> </thead> <tbody> <tr> <td>0</td> <td>21.5 (Estimate)</td> </tr> <tr> <td>5</td> <td>18.5</td> </tr> <tr> <td>10</td> <td>16.0</td> </tr> <tr> <td>15</td> <td>14.0</td> </tr> <tr> <td>20</td> <td>11.5</td> </tr> <tr> <td>25</td> <td>9.0</td> </tr> </tbody> </table>	Time / minutes	Volume of sodium bisulfate / cm ³	0	21.5 (Estimate)	5	18.5	10	16.0	15	14.0	20	11.5	25	9.0	Just 'V/cm ³ ' T for time	3
Time / minutes	Volume of sodium bisulfate / cm ³																
0	21.5 (Estimate)																
5	18.5																
10	16.0																
15	14.0																
20	11.5																
25	9.0																

Question Number	Acceptable Answers	Reject	Mark
2(b)(iii)	(Gradient numerical value) = (-) 0.48 to 0.52 IGNORE SF (1)		2
	(Gradient units) = $\text{cm}^3 \text{min}^{-1}$ ALLOW cm^3 / min (1)		
	ALLOW (-) $8.33 \times 10^{-3} \text{cm}^3 \text{s}^{-1}$ for two marks		

Question Number	Acceptable Answers	Reject	Mark
2(c)	O / zero / zeroth (order) and because the graph is a straight line OR Rate/gradient is constant OR rate doesn't depend on iodine concentration/volume	Positive gradient / Concentration is proportional to time	1

Question Number	Acceptable Answers	Reject	Mark
2(d)	If the time is known (accurately) then it can still be plotted correctly/The actual time doesn't matter as long as it is known (accurately)		1

Question Number	Acceptable Answers	Reject	Mark
2(e)	There is insufficient volume of the reaction mixture left in the flask (to pipette exactly 10.0cm^3) Ignore there is only 60cm^3 of reaction mixture (1) EITHER Add sodium hydrogencarbonate directly to the flask with the reaction mixture (and then titrate the mixture)/ Titrate the remaining reaction mixture (in the flask) ALLOW Pour the (remaining) reaction mixture into the sodium hydrogencarbonate solution OR Use a 5cm^3 pipette and double the titre value (1)	References to 8.50cm^3	2

Question Number	Acceptable Answers	Reject	Mark
2(f)	Temperature (1) Water bath ALLOW Other suitable lab equipment which would control the temperature/ice-water mixture/oil bath No TE on other factors Ignore references to thermometers/ thermostatically controlled rooms/air-conditioning (1)	Heat	2

(Total for Question 2 = 13 marks)

Question Number	Acceptable Answers	Reject	Mark
3(a)(i)	Results are not concordant/not within 0.1(0) cm ³ / not within 0.2(0) cm ³		1

Question Number	Acceptable Answers	Reject	Mark
3(a)(ii)	Colourless to (pale) pink ALLOW Yellowy brown for starting colour	Clear Red Purple Magenta	1

Question Number	Acceptable Answers	Reject	Mark
3(a)(iii)	The colour of the apple juice (will make the colour change at the end-point difficult to determine) ALLOW Cloudiness/Not clear / any sensible colour for apple juices such as green or brown		1

Question Number	Acceptable Answers	Reject	Mark
3(a)(iv)	<p>Method 1</p> <p>Number of moles of OH⁻ =(0.100 x 0.01680=) 1.68 x 10⁻³ (mol) (1)</p> <p>Number of moles of malic acid = (1.68 x 10⁻³ ÷ 2 =) 8.40 x 10⁻⁴ (mol) (1)</p> <p>Concentration of malic acid = (8.40 x 10⁻⁴ ÷ 0.025= = 3.36 x 10⁻² / 0.0336 =) = 3.4 x 10⁻² / 0.034 (mol dm⁻³) (1)</p> <p>OR</p> <p>Method 2</p> <p>Step 1: Volume ratio 16.8/25 (1)</p> <p>Step 2: Multiply by 0.100 (1)</p> <p>Step 3: Multiply by ½ to give 3.4 x 10⁻² / 0.034 (mol dm⁻³) (1)</p> <p>Correct answer to 2SF without working scores (3)</p> <p>If units given, then they must be correct</p> <p>ALLOW TE from each step of the calculation</p>	<p>Any answer not to 2 SF</p> <p>Any answer not to 2 SF</p> <p>mol/dm⁻³</p>	3

Question Number	Acceptable Answers	Reject	Mark
3(a)(v)	<p>The mean titre would be 25.2(0) cm³ / 2.52 x 10⁻² dm³ (1)</p> <p>There are three (carboxylic) acid groups in isocitric acid and compared to two in malic acid (1)</p> <p>ALLOW for 1 mark the titre would be greater because there is one more (carboxylic) acid groups in isocitric acid</p> <p>ALLOW for 1 mark the titre would be $\frac{3}{2}$ greater if no other mark awarded</p>	<p>One more OH group</p>	2

Question Number	Acceptable Answers	Reject	Mark
3(b)(i)	Triplet / three splits / split into three ALLOW Just '3' / 1:2:1 / triple and the adjacent carbon has two hydrogen atoms but ignore just (n+1) rule unless explained ALLOW and there are two adjacent hydrogen atoms		1

Question Number	Acceptable Answers	Reject	Mark
3(b)(ii)	Two/2 (peaks)		1

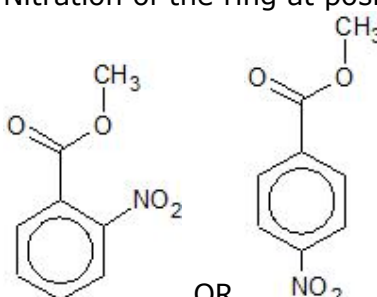
Question Number	Acceptable Answers	Reject	Mark
3(b)(iii)	(CH ₃) ₄ Si/Tetramethylsilane ALLOW TMS If formula and name given, then both must be correct	SiCl ₄ TMC	1

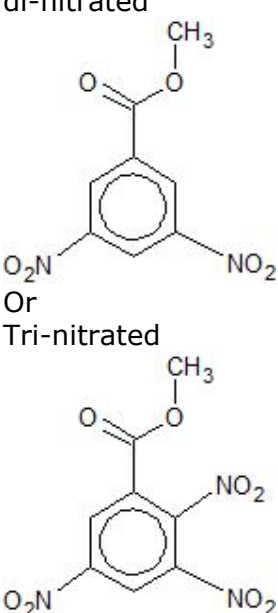
Question Number	Acceptable Answers	Reject	Mark
3(c)(i)	COOH ⁺ /CO ₂ H ⁺	Just 'displayed formula'	1

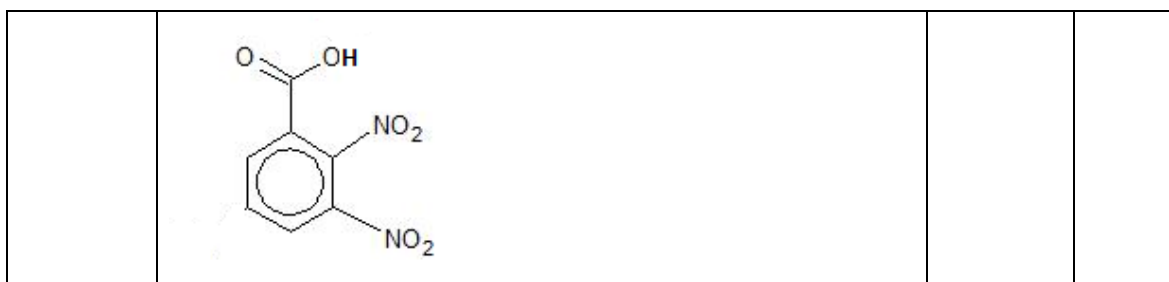
Question Number	Acceptable Answers	Reject	Mark
3(c)(ii)	C ₃ H ₅ O ₃ ⁺ ALLOW the atoms in any order	Just 'structural/ displayed/ skeletal formulae'	1

(Total for Question 3 = 13 marks)

Question Number	Acceptable Answers	Reject	Mark
4(a)	Ice cubes have limited/less surface area/less touching area/ less contact for cooling Accept reverse argument IGNORE Cooling is more efficient/cooling is faster		1

Question Number	Acceptable Answers	Reject	Mark
4(b)(i)	Nitration of the ring at position 2 or 4 	Nitration of any other part of the molecule Multi-nitrations	1

Question Number	Acceptable Answers	Reject	Mark
4(b)(ii)	Any further-substituted product, such as di-nitrated  Or Tri-nitrated Accept substitution at any or all of the five available positions ALLOW Hydrolysed ester, eg	Nitration of methyl group	1



Question Number	Acceptable Answers	Reject	Mark
4(c)	<p>Any two from (Modification 1) perforations / holes in the Buchner funnel need to be shown (1)</p> <p>(Modification 2) One of the tubes from the tap should go straight down (to create suction) and a second tube should go to the sidearm of the Buchner flask</p> <p>ALLOW Replace attachment to tap with to vacuum/ pump/suction pump (1)</p> <p>(Modification 3) Filter paper should be trimmed so that it does not go up at the sides of the funnel (1)</p> <p>Diagrams of modifications alone or to support descriptions can be awarded credit</p>		2

Question Number	Acceptable Answers	Reject	Mark
4(d)	<p>M1 Dissolve in the minimum/small volume of hot/boiling ethanol / solvent</p> <p>ALLOW Add ethanol and dissolve, heat to evaporate some ethanol to produce a saturated solution / crystals form on the end of a glass rod (1)</p> <p>Ignore reference to hot filtration</p> <p>M2 Cool in an ice bath (to form the crystals) ALLOW Leave to cool/Allow to cool (1)</p> <p>M3 Filter using vacuum filtration/suction filtration/ Buchner funnel (to remove soluble impurities) Ignore 'wash with cold solvent' (1)</p> <p>M4 Dry (crystals) between two pieces of filter paper/ (pat) dry with filter paper in a desiccator/ in a cool or warm oven (1)</p>	<p>Use of water as solvent/</p> <p>Just 'dissolve and then heat'</p> <p>Wash with (cold) water</p> <p>Just 'oven' / Add a drying agent</p>	4

Question Number	Acceptable Answers	Reject	Mark
4(e)	<p>$n(\text{methyl benzoate}) = 2.0 \div 136 = 0.0147.. \text{ (mol)}$ (1)</p> <p>$= n(\text{methyl 3-nitrobenzoate})$</p> <p>$m(\text{methyl 3-nitrobenzoate}) = 0.0147.. \times 181 = 2.66.. \text{ (g)}$</p> <p>so $73\% = 2.66.. \times 0.73 = 1.943/1.94/1.9 \text{ (g)}$ (1)</p> <p>Correct answer without working scores (2) Ignore SF except 1</p>	<p>Intermediate rounding to 1SF</p>	2

Question Number	Acceptable Answers	Reject	Mark
4(f)	<p>The range would start at a lower temperature / OR</p> <p>The range would be wider/larger</p> <p>ALLOW Range given as numbers provided the highest temperature stated is not $>80^{\circ}\text{C}$ and the lowest is not $<70^{\circ}\text{C}$</p>	<p>Temperatures would be raised Just 'wide'/ 'large'</p>	1

(Total for Question 4 = 12 marks)

Total for Paper = 50 marks

