



Mark Scheme (Results)

Summer 2018

Pearson Edexcel International Advance
Level In Chemistry (WCH03) Paper 01
Chemistry Laboratory Skill I

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
 - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
 - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
 - iii) organise information clearly and coherently, using specialist vocabulary when appropriate

Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

() means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer.

ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

Quality of Written Communication

Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

Full marks will be awarded if the candidate has demonstrated the above abilities. Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------------|
| 1(a)(i) | <p>Cations Any two from: (1) mark each Calcium / Ca^{2+} / Strontium / Sr^{2+} / Barium / Ba^{2+}</p> <p>ALLOW Charges written as +2 (2)</p> <p>IGNORE Radium / Ra^{2+}</p> <p>Max (1) if additional incorrect species</p> <p>Anion Nitrate(V) / NO_3^-</p> <p>ALLOW nitrate (1)</p> | <p>For all marks in (a):</p> <p>If formulae are given, penalise incorrect / missing charges once only in (a)(i) or (a)(iii)</p> <p>If name and formulae are given both must be correct, but just penalise one incorrect answer once only in (a)(i) to (a)(iii)</p> | (3) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--------|------------|
| 1(a)(ii) | <p>(Gas C) nitrogen dioxide / nitrogen(IV) oxide / NO_2</p> <p>ALLOW dinitrogen tetroxide / N_2O_4 (1)</p> <p>(Gas D) oxygen / O_2 (1)</p> <p>ALLOW (1) for two correct gases in the wrong order</p> | | (2) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|---|------|
| 1(a)(iii) | <p>Mark independently</p> <p>Flame test</p> <p>ALLOW Description of a flame test (1)</p> <p>Any two from: (1) mark each Calcium / Ca²⁺: brick-red / orange-red / yellow-red</p> <p>ALLOW Just 'red' for calcium if crimson or scarlet used for strontium or if strontium is not identified</p> <p>Strontium / Sr²⁺: red / crimson /scarlet</p> <p>Barium / Ba²⁺: (apple) green (2)</p> <p>Note If other Group 2 cations identified in (i):</p> <p>ALLOW Magnesium / Mg²⁺ / beryllium / Be²⁺: no colour Radium / Ra²⁺: carmine / red</p> <p>IGNORE Shades of colours e.g. pale / bright</p> <p>No TE on ions from other groups</p> | <p>Just 'orange'</p> <p>Yellow-green / blue-green</p> | (3) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------------|
| 1(b)(i) | Anion B : Br ⁻ IGNORE bromide | Br / Br ₂ / Br ₂ ⁻ just 'bromine' Iodide / I ⁻ / I | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--|------------|
| 1(b)(ii) | Gas E : Bromine / Br ₂ | If name and formulae are given penalise one incorrect answer once only in (b)(ii) to (iv) Bromide / Br ⁻ / Br Iodine / I ₂ Nitrogen dioxide / NO ₂ | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|------------------|---|-------------------------|------------|
| 1(b)(iii) | Hydrogen bromide / HBr Note If iodide identified in (i), ALLOW hydrogen iodide / HI | Hydrogen chloride / HCl | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|-----------------|------------|
| 1(b)(iv) | Sulfur dioxide / sulfur(IV) oxide / SO ₂ Note If iodide identified in (i), ALLOW hydrogen sulfide / H ₂ S | Sulphurous acid | (1) |

(Total for Question 1 = 12 marks)

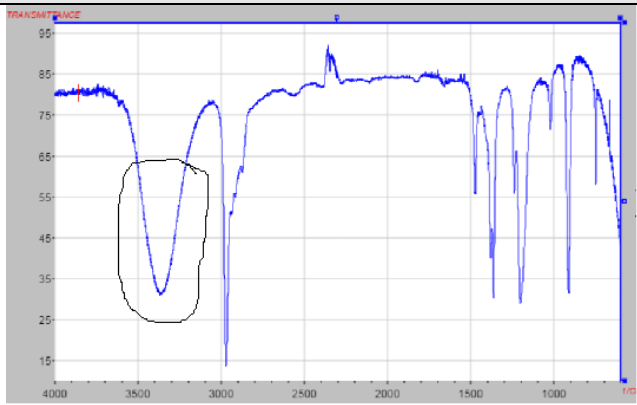
| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|---|------------|
| 2(a)(i) | Alcohol ALLOW –OH / OH / O–H / –O–H / hydroxyl / hydroxy IGNORE Alkyl / butanol | OH ⁻ / hydroxide / hydroxylic / carboxylic acid / any other functional group | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|---|------------|
| 2(a)(ii) | $ \begin{array}{cccc} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array} $ <p style="text-align: right;">(1)</p> $ \begin{array}{cccc} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{H} \\ & & & \\ \text{H} & \text{H} & \text{OH} & \text{H} \end{array} $ <p style="text-align: right;">(1)</p> $ \begin{array}{cccc} \text{H} & \text{CH}_3 & \text{H} & \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{OH} \\ & & & \\ \text{H} & \text{H} & \text{H} & \end{array} $ <p style="text-align: right;">(1)</p> $ \begin{array}{cccc} \text{H} & \text{CH}_3 & \text{H} & \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{H} \\ & & & \\ \text{H} & \text{OH} & \text{H} & \end{array} $ <p style="text-align: right;">(1)</p> <p>ALLOW Alcohols in any order</p> <p>ALLOW Displayed or structural formulae or any combination of these / skeletal formulae but if more than one type of structure is given, all must be correct</p> <p>IGNORE C–OH connectivity with vertical bonds</p> <p>IGNORE Names even if incorrect</p> | <p>Penalise OH–C on left side of structure / C–HO on right side of structure once only in (a)(ii) and (c)</p> <p>Penalise missing H in C–H bonds once only in (a)(ii) and (c)</p> <p>Penalise missing bond once only in (a)(ii) and (c)</p> | (4) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--------|------------|
| 2(b) | Tertiary (alcohol / hydroxyl) ALLOW 3° / 3^y (alcohol / hydroxyl) IGNORE Not primary or secondary alcohol / the OH is attached to a carbon that is attached to 3 other carbon atoms | | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|---|------------|
| 2(c) | Stand alone mark $ \begin{array}{ccccc} & \text{H} & & \text{CH}_3 & \text{H} \\ & & & & \\ \text{H} & - \text{C} & - & \text{C} & - \text{C} & - \text{H} \\ & & & & \\ & \text{H} & & \text{OH} & \text{H} \end{array} $ ALLOW Displayed or structural formulae or any combination of these / skeletal formulae ALLOW (2-)methylpropan-2-ol ALLOW Any reference to correct structure drawn in (a)(ii) IGNORE C–OH connectivity with vertical bonds No TE on wrong type of alcohol identified in (b) | Penalise OH–C on left side of structure / C–HO on right side of structure once only in (a)(ii) and (c) | (1) |

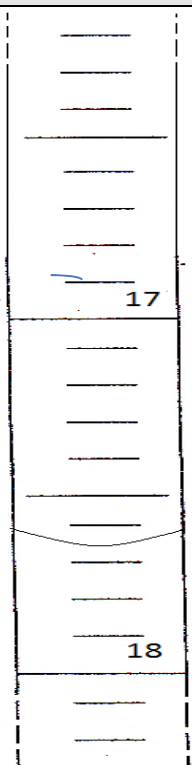
| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|-------------------|------------|
| 2(d)(i) | $(\text{CH}_3)_3\text{COH} + \text{PCl}_5 \rightarrow (\text{CH}_3)_3\text{CCl} + \text{HCl} + \text{POCl}_3$ ALLOW Displayed, structural or skeletal formulae or any combination of these ALLOW $\text{C}(\text{CH}_3)_3\text{OH} / \text{PCl}_3\text{O}$ ALLOW $\text{C}_4\text{H}_9\text{OH} + \text{PCl}_5 \rightarrow \text{C}_4\text{H}_9\text{Cl} + \text{HCl} + \text{POCl}_3$ OR $\text{C}_4\text{H}_{10}\text{O} + \text{PCl}_5 \rightarrow \text{C}_4\text{H}_9\text{Cl} + \text{HCl} + \text{POCl}_3$ TE on alcohol in 2(c) IGNORE State symbols even if incorrect | Equation with ROH | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------------|
| 2(d)(ii) |  <p>Correct absorbance circled</p> <p>ALLOW Other ways of identifying the absorbance e.g. label with OH</p> <p>ALLOW All or any part of the absorbance circled</p> | Additional absorbances circled or labelled | (1) |

(Total for Question 2 = 9 marks)

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--|------|
| 3(a)(i) | <p>Pipette Use of a (25.0 cm³) pipette (and pipette filler) to transfer sodium hydroxide to a conical flask</p> <p>ALLOW Beaker / flask for conical flask (1)</p> <p>Mixing (Add acid to the alkali while) swirling / mixing</p> <p>ALLOW Stirring / shaking (1)</p> <p>Dropwise Add (sulfuric) acid drop-by-drop / slowly (until indicator just changes colour) (1)</p> <p>IGNORE Slowly open the tap</p> <p>End-point Use of methyl orange/ indicator and colour change yellow to peach/orange/ (and read the burette volume) (1)</p> <p>IGNORE Red / pink</p> <p>Technique - Any one from: Rinse flask with distilled / deionised water</p> <p>OR Touch tip of pipette onto surface of solution</p> <p>OR Use of white tile / paper</p> <p>OR Add (sulfuric) acid to the sodium hydroxide / alkali / conical flask until a specified volume between 15 and 17.5 cm³ is added (before adding dropwise)</p> <p>OR Wash drops of solution into flask with distilled / deionised water (1)</p> | <p>Volumetric / standard flask</p> <p>Use of any other indicator</p> <p>Wash flask with NaOH negates any point from the technique mark</p> | (5) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--------|------|
| 3(a)(ii) | Boxes ticked under 17.55 and 17.65 cm ³ and mean titre is 17.60 (cm ³) ALLOW 17.6 (cm ³) | | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--------|------|
| 3(a)(iii) |  <p>Mark independently Bottom of meniscus anywhere between 17.6 and 17.7 (cm³)</p> <p>ALLOW Any indication of value between 17.6 and 17.7 (cm³) e.g. an arrow (1)</p> <p>Shape of meniscus</p> <p>ALLOW Any values for top of curve (1)</p> | | (2) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|---|------------|
| 3(b) | <p>Stand alone marks</p> <p>First mark Repeat the titration using the mean titre but no indicator OR Add 17.6(0) cm³ / mean titre of (sulfuric) acid to 25(.0) cm³ of sodium hydroxide / alkali (with no mention of indicator added)</p> <p>ALLOW Quoted value of 17.7 cm³ (1)</p> <p>Second mark (Transfer the mixture to an evaporating basin and) heat the solution to evaporate some / most of the water OR Heat to crystallisation point / saturation point / concentrate the solution OR Leave the solution for the water to evaporate</p> <p>ALLOW Heat the solution until (some) crystals form (1)</p> <p>IGNORE Any reference to filtering / washing crystals / drying by any method</p> | <p>Evaporate all water / to dryness</p> | (2) |

(Total for Question 3 = 10 marks)

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|--|------------|
| 4(c)(i) | (amount of anhydrous zinc sulfate left) $= \frac{2.04}{161.5}$ $= 0.012632 / 0.01263 / 0.0126 / 0.013$ (mol) ALLOW 0.0127 from use of 161 TE on mass of contents after heating IGNORE SF except 1SF Correct answer with no working scores (1) | Penalise 1SF in (c)(i) and (c)(ii) once only | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------------|
| 4(c)(ii) | (amount of water lost) $= \frac{1.41}{18}$ $= 0.078333/0.07833/0.0783/0.078$ (mol) TE on mass of water lost IGNORE SF except 1SF Correct answer with no working scores (1) | Penalise 1SF in (c)(i) and (c)(ii) once only | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|------------------|---|--------|------------|
| 4(c)(iii) | (Value of n) $= \frac{0.078333}{0.012632}$ $= 6.2014/6.201/6.20/6.2/6$ ALLOW other answers based on rounded values used in (c)(i) and (c)(ii) TE on (c)(i) and (c)(ii) IGNORE SF, including 1 SF Correct answer with no working scores (1) | n = 7 | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--------|------------|
| 4(d)(i) | Error 1 (Effect on measured mass of water lost) increases / more and (Effect on value of n) increases / more (1) Error 2 (Effect on measured mass of water lost) decreases / less and (Effect on value of n) decreases / less (1) | | (2) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------------|
| 4(d)(ii) | Put a lid on the crucible / use a lid OR Heat more gently / slowly (at start followed by stronger heating) ALLOW Use of a cover (for the crucible) IGNORE Electrical heating / Lift the lid (at intervals) / Use larger crucible / larger crystals / Use fewer crystals / Just 'add a top' Small holes in the lid | Use a stopper on the crucible / seal the crucible Use an evaporating basin / boiling tube / beaker / any other container Add anti-bumping granules | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|------------------|---|--------------------|------------|
| 4(d)(iii) | Heat to constant mass OR Description of heating to constant mass ALLOW Heat to constant weight / do the reaction to constant mass IGNORE Just 'until measurements are the same' Heat for a long time / Use of cobalt chloride paper / anhydrous copper sulfate (to test for water vapour) | Add a drying agent | (1) |

(Total for Question 4 = 10 marks)

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--------|------------|
| 5(a) | <p>Any one from: Promotes smooth boiling/uniform heating</p> <p>OR</p> <p>Prevents localised / uneven / uncontrolled / violent / flash boiling / heating</p> <p>OR</p> <p>Promotes the formation of small bubbles</p> <p>OR</p> <p>Prevents the formation of large bubbles</p> <p>OR</p> <p>Provides nucleation centres / rough surface (for bubble formation)</p> <p>OR</p> <p>Prevents superheating</p> <p>IGNORE</p> <p>Prevents bumping / spitting / explosion / liquid splashing out So the reaction occurs smoothly Prevents vigorous reaction / boiling</p> | | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|------------------------------------|------------|
| 5(b) | <p>An exothermic reaction occurs / reaction releases heat (energy)</p> <p>IGNORE</p> <p>Vigorous/violent reaction / Just 'the temperature increases during the reaction / addition' / Reagents are volatile / Sulfuric acid is corrosive</p> | Reference to explosion / explosive | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|---|------|
| 5(c) | <p>Ethanol / alcohol is (highly) flammable</p> <p>ALLOW Ethyl ethanoate/ ethanoic acid / organic compounds (highly) flammable</p> <p>OR The flask contains some flammable liquids</p> <p>OR To keep the temperature of the flask below 100°C</p> <p>IGNORE To ensure even heating / Just 'it is volatile' Just 'the reactants / products / contents are flammable' / Just 'it / the solution / liquid / mixture is flammable'</p> | <p>Sulfuric acid is flammable</p> <p>Reference to explosion</p> | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|---|---|------|
| 5(d) | <p>To prevent the loss / escape of any volatile substances / volatile reactants / volatile products / organic compounds / named organic compound</p> <p>OR To make sure that vapour(s) condense / to prevent vapour escaping</p> <p>ALLOW Gas for vapour</p> <p>ALLOW To increase the yield</p> <p>ALLOW So that the reaction goes to completion</p> <p>IGNORE Answers related to safety / flammability / slow rate of reaction</p> | <p>To prevent evaporation / boiling</p> <p>Any reference to oxidation</p> | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--------|------|
| 5(e)(i) | Carbon dioxide / CO ₂ (gas / (g)) | | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|--|------|
| 5(e)(ii) | Hold separating funnel upside down and open the tap OR Remove / open the stopper / bung / cork / lid / top of the funnel(from time to time) IGNORE Use of finger instead of stopper | Open the tap without inverting the funnel Remove the stopper while the funnel is inverted | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|-------------------|------|
| 5(f)(i) | As a drying agent / to dry it ALLOW To remove / absorb water / moisture IGNORE To absorb (unreacted) ethanol | Dehydrating agent | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|---|------|
| 5(f)(ii) | (Anhydrous) sodium sulfate / Na ₂ SO ₄ / magnesium sulfate / MgSO ₄ / calcium sulfate / CaSO ₄ ALLOW Silica gel If name and formula are given, both must be correct IGNORE Mention of dehydrating agent | Copper(II) sulfate Cobalt chloride Sodium / potassium carbonate / Sulfuric acid | (1) |

| Question Number | Acceptable Answers | Reject | Mark |
|-----------------|--|-------------------|------------|
| 5(g) | Lower number 74 or 75 or 76(°C) and Higher number 78 or 79 or 80(°C) ALLOW numbers in reverse order e.g. 79-75(°C) | Any single number | (1) |

(Total for Question 5 = 9 marks)

TOTAL FOR PAPER = 50 MARKS