



Pearson

Mark Scheme (Results)

Summer 2017

Pearson Edexcel IAL
In Chemistry (WCH05) Paper 01
General Principles of Chemistry II –
Transition Metals and Organic
Nitrogen Chemistry

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
 - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
 - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
 - iii) organise information clearly and coherently, using specialist vocabulary when appropriate

Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

() means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer.

ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

Quality of Written Communication

Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

Full marks will be awarded if the candidate has demonstrated the above abilities.

Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

Section A (multiple choice)

Question Number	Answer	Mark
(1)	<p>The only correct answer is C</p> <p><i>A is not correct because these are in the d-block but are not transition metals</i></p> <p><i>B is not correct because tin is in Group 4</i></p> <p><i>D is not correct because these are in the d-block but are not transition metals</i></p>	(1)

Question Number	Incorrect answers	Mark
(2)	<p>The only correct answer is A</p> <p><i>B is not correct because there should be a decrease of 2 oxidation numbers as the ratio of $Tl^{3+} : I^- = 1 : 2$</i></p> <p><i>C is not correct because there should be a decrease in oxidation number as iodide ions are oxidised so thallium ions are reduced</i></p> <p><i>D is not correct because there should be a decrease in oxidation number as iodide ions are oxidised so thallium ions are reduced</i></p>	(1)

Question Number	Incorrect answers	Mark
(3)	<p>The only correct answer is C</p> <p><i>A is not correct because light is not emitted when an electron drops back to the ground state</i></p> <p><i>B is not correct because this happens in a flame test</i></p> <p><i>D is not correct because light is not emitted when an electron is promoted</i></p>	(1)

Question Number	Incorrect answers	Mark
(4)	<p>The only correct answer is C</p> <p><i>A is not correct because do not give a pale blue precipitate with aqueous copper(II) sulfate</i></p> <p><i>B is not correct because do not give a pale blue precipitate with aqueous copper(II) sulfate</i></p> <p><i>D is not correct because do not give a pale blue precipitate with aqueous copper(II) sulfate</i></p>	(1)

Question Number	Incorrect answers	Mark
(5)	<p>The only correct answer is A</p> <p><i>B is not correct because uses MnO_4^- concentration as $0.0100 \text{ mol dm}^{-3}$</i></p> <p><i>C is not correct because uses mole ratio the wrong way round</i></p> <p><i>D is not correct because uses mole ratio the wrong way round and MnO_4^- concentration as $0.0100 \text{ mol dm}^{-3}$</i></p>	(1)

Question Number	Incorrect answers	Mark
(6)	<p>The only correct answer is A</p> <p><i>B is not correct because this is the oxidation number of C in $C_2O_4^{2-}$</i></p> <p><i>C is not correct because this is the oxidation number of C in CO_2</i></p> <p><i>D is not correct because this is the change in oxidation number of one Mn</i></p>	(1)

Question Number	Incorrect answers	Mark
(7)	<p>The only correct answer is D</p> <p><i>A is not correct because would form a precipitate of silver carbonate</i></p> <p><i>B is not correct because would form a precipitate of silver chloride</i></p> <p><i>C is not correct because would form a precipitate of silver iodide</i></p>	(1)

Question Number	Incorrect answers	Mark
(8)	<p>The only correct answer is C</p> <p><i>A is not correct because do not form hydroxide ions which are alkaline and turn phenolphthalein pink</i></p> <p><i>B is not correct because do not form hydroxide ions which are alkaline and turn phenolphthalein pink</i></p> <p><i>D is not correct because do not form hydroxide ions which are alkaline and turn phenolphthalein pink</i></p>	(1)

Question Number	Incorrect answers	Mark
9(a)	<p>The only correct answer is A</p> <p><i>B is not correct because will only reduce chlorine</i></p> <p><i>C is not correct because will not reduce anything in that list</i></p> <p><i>D is not correct because this is the strongest oxidising agent</i></p>	(1)

Question Number	Incorrect answers	Mark
9(b)	<p>The only correct answer is D</p> <p><i>A is not correct because iodine will also react in this way</i></p> <p><i>B is not correct because bromine will also react in this way</i></p> <p><i>C is not correct because chlorine oxidises chromium(III) to chromium(VI)</i></p>	(1)

Question Number	Incorrect answers	Mark
(10)	<p>The only correct answer is C</p> <p><i>A is not correct because are provided by X-ray diffraction</i></p> <p><i>B is not correct because are provided by X-ray diffraction</i></p> <p><i>D is not correct because are provided by X-ray diffraction</i></p>	(1)

Question Number	Incorrect answers	Mark
(11)	<p>The only correct answer is D</p> <p><i>A is not correct because incorrect products</i></p> <p><i>B is not correct because incorrect products</i></p> <p><i>C is not correct because incorrect products</i></p>	(1)

Question Number	Incorrect answers	Mark
(12)	<p>The only correct answer is A</p> <p><i>B is not correct because nucleophile is incorrect</i></p> <p><i>C is not correct because base is incorrect</i></p> <p><i>D is not correct because base and electrophile are incorrect</i></p>	(1)

Question Number	Incorrect answers	Mark
(13)	<p>The only correct answer is B</p> <p><i>A is not correct because the 3 carbon atoms in the middle of the alkyl chain are chiral</i></p> <p><i>C is not correct because the 3 carbon atoms in the middle of the alkyl chain are chiral</i></p> <p><i>D is not correct because the 3 carbon atoms in the middle of the alkyl chain are chiral</i></p>	(1)

Question Number	Incorrect answers	Mark
(14)	<p>The only correct answer is B</p> <p><i>A is not correct because this is 70% of the moles of nitrobenzene</i></p> <p><i>C is not correct because this is 70% of 2.46 g</i></p> <p><i>D is not correct because the M_rs have been mixed up</i></p>	(1)

Question Number	Incorrect answers	Mark
(15)	<p>The only correct answer is C</p> <p><i>A is not correct because both phenol groups react with bromine water</i></p> <p><i>B is not correct because both amine groups react with copper(II) sulfate solution</i></p> <p><i>D is not correct because neither molecule has an aldehyde group to react with Tollens' reagent</i></p>	(1)

Question Number	Incorrect answers	Mark
(16)	<p>The only correct answer is D</p> <p><i>A is not correct because would not give the absorbance due to OH at 3300 to 2500 cm⁻¹</i></p> <p><i>B is not correct because would not give the absorbance due to OH at 3300 to 2500 cm⁻¹</i></p> <p><i>C is not correct because would give OH absorbance at 3750 to 3200 cm⁻¹</i></p>	(1)

Question Number	Incorrect answers	Mark
(17)	<p>The only correct answer is B</p> <p><i>A is not correct because the amine groups should be in positions 3, 5 relative to the methyl group, not 2,6</i></p> <p><i>C is not correct because the amine groups should be in positions 3, 5 relative to the methyl group, not 2,3</i></p> <p><i>D is not correct because the amine groups should be in positions 3, 5 relative to the methyl group, not 2,5</i></p>	(1)

Question Number	Incorrect answers	Mark
(18)	<p>The only correct answer is C</p> <p>A is not correct because the potassium salt of the carboxylic acid should be formed</p> <p>B is not correct because the potassium salt of the carboxylic acid should be formed and the potassium salt of the alcohol does not form</p> <p>D is not correct because the potassium salt of the alcohol does not form</p>	(1)

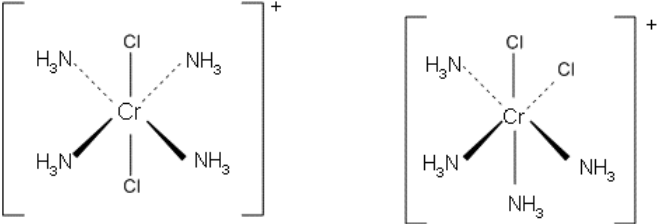
Question Number	Incorrect answers	Mark
(19)	<p>The only correct answer is B</p> <p>A is not correct because there are 6 amino acids in the structure but 1st, 3rd and 5th are the same and 4th and 6th are the same so only 3 different amino acids</p> <p>C is not correct because there are 6 amino acids in the structure but 1st, 3rd and 5th are the same and 4th and 6th are the same so only 3 different amino acids</p> <p>D is not correct because there are 6 amino acids in the structure but 1st, 3rd and 5th are the same and 4th and 6th are the same so only 3 different amino acids</p>	(1)

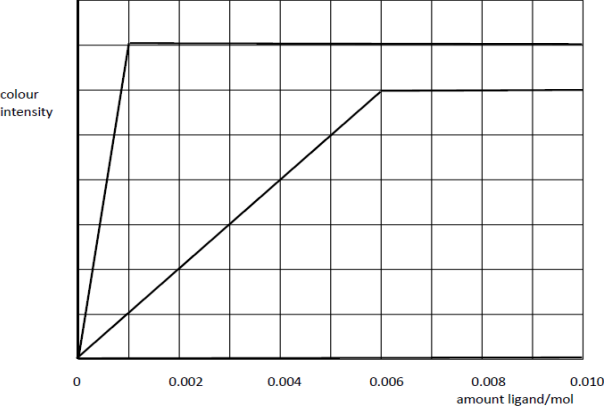
Total for Section A = 20 MARKS

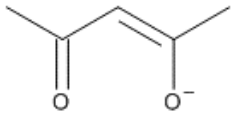
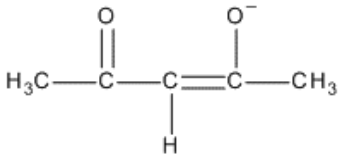
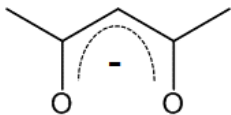
Section B

Question Number	Acceptable Answers	Reject	Mark
20(a)	<p>Shape – tetrahedral (1)</p> <p>Explanation – (4) pairs of (valence) electrons / (4) bond pairs and arranged to minimise repulsion</p> <p>ALLOW (4) pairs of (valence) electrons / (4) bond pairs and arranged with maximum separation (1)</p> <p>ALLOW Only TE on square planar with a description of either 4 bonding pairs or 4 bonding pairs and 2 lone pairs and maximum separation/minimum repulsion</p> <p>IGNORE incorrect bond angle</p>	<p>Just 'electrons repel'</p> <p>Just '(Dative) bonds' / atoms/ ligands/ Cl repel</p>	(2)

Question Number	Acceptable Answers	Reject	Mark
20(b)	<p>Forming precipitate</p> $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s}) + 6\text{H}_2\text{O}(\text{l})$ <p>OR</p> $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{H}_2\text{O})_3(\text{s}) + 3\text{H}_2\text{O}(\text{l})$ <p>ALLOW</p> $\text{Cr}^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s}) \quad \mathbf{(1)}$ <p>Dissolving precipitate</p> $\text{Cr}(\text{OH})_3(\text{s}) + 3\text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_6]^{3-}(\text{aq})$ <p>OR</p> $\text{Cr}(\text{OH})_3(\text{H}_2\text{O})_3(\text{s}) + 3\text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_6]^{3-}(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$ <p>ALLOW</p> $\text{Cr}(\text{OH})_3(\text{s}) + \text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_4]^{-}(\text{aq})$ <p>ALLOW</p> $\text{Cr}(\text{OH})_3(\text{H}_2\text{O})_3(\text{s}) + \text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_4]^{-}(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$ <p>ALLOW</p> $\text{Cr}(\text{OH})_3(\text{H}_2\text{O})_3(\text{s}) + \text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_4(\text{H}_2\text{O})_2]^{-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \quad \mathbf{(1)}$ <p>Notes</p> <p>ALLOW (1) for two correct non-ionic equations with NaOH / Na⁺ + OH⁻</p> <p>ALLOW (1) for two unbalanced equations with correct species and state symbols</p> <p>IGNORE square brackets around neutral species</p> <p>IGNORE the order of ligands in formulae with OH and H₂O</p> <p>IGNORE charges inside the brackets</p> <p>If no other mark is awarded, allow (1) for</p> $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 6\text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_6]^{3-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$ <p>OR</p> $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 4\text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_4]^{-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$ <p>OR</p> $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 4\text{OH}^{-}(\text{aq}) \rightarrow [\text{Cr}(\text{OH})_4(\text{H}_2\text{O})_2]^{-}(\text{aq}) + 4\text{H}_2\text{O}(\text{l})$	Incorrect or missing state symbols once only	(2)

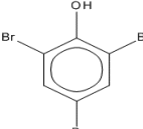
Question Number	Acceptable Answers	Reject	Mark
20(c)	 <p>Any trans isomer (1)</p> <p>Any cis isomer (1)</p> <p>IGNORE connectivity of Cr-NH₃</p> <p>IGNORE charges on Cr and Cl</p>	<p>Cl₂ once only</p> <p>NH₄ once only</p> <p>additional isomer once only</p>	(2)

Question Number	Acceptable Answers	Reject	Mark
20(d)	 <p>Mark independently</p> <p>A rise to 0.001 mol ligand (1)</p> <p>Horizontal line of any length with any value for colour intensity greater than original line (1)</p>		(2)

Question Number	Acceptable Answers	Reject	Mark
20(e)	 <p>ALLOW structural, displayed or skeletal formulae or any combination of these e.g.</p>  <p>ALLOW delocalised structure</p>  <p>ALLOW structure in brackets with charge outside</p> <p>IGNORE lone pairs</p> <p>IGNORE additional structures as working</p>	<p>Charge on double bonded oxygen</p> <p>Charge on both oxygens</p>	(1)

(Total for Question 20 = 9 marks)

Question Number	Acceptable Answers	Reject	Mark
21(a)(i)	<p>White precipitate forms</p> <p>ALLOW solid / crystals / ppt for precipitate</p> <p>IGNORE antiseptic smell / colour change</p> <p>IGNORE cloudy</p> <p>IGNORE name of ppt even if incorrect</p>	Incorrect colour of ppt	(1)

Question Number	Acceptable Answers	Reject	Mark
21(a)(ii)	<p>$C_6H_5OH + 3Br_2 \rightarrow 3HBr +$</p> <div style="text-align: center;">  </div> <p>First mark - organic product</p> <p>ALLOW Kekule structure</p> <p>ALLOW substitution of Br to any 3 positions on the ring / $C_6H_2(Br)_3OH$ / $C_6H_2Br_3OH$ / $C_6H_2(OH)Br_3$ / $C_6H_2OHB r_3$ (1)</p> <p>IGNORE connectivity to OH</p> <p>IGNORE name even if incorrect</p> <p>Second mark Rest of equation correct Phenol may be drawn (1)</p> <p>Note Mono or di substitution scores (1) for balanced equation</p>	Molecular formula e.g. $C_6H_3OBr_3$	(2)

Question Number	Acceptable Answers	Reject	Mark
*21(a)(iii)	<p>First mark The lone pair (of electrons) on the O (of OH)</p> <p>and</p> <p>EITHER Overlaps with the π/delocalised electrons in the benzene ring / delocalised system</p> <p>OR Feeds into / donates into / interacts with (benzene) ring / delocalised electrons / delocalised system</p> <p>ALLOW Increases the electron density of the (benzene) ring (1)</p> <p>Second mark (Increased electron density) makes the ring more susceptible to electrophilic attack / attack by Br^+ / $\text{Br}^{\delta+}$</p> <p>ALLOW Phenol is a better nucleophile (1)</p>	Ring is more electronegative	(2)

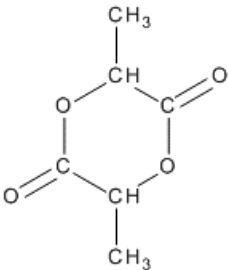
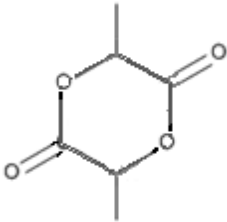
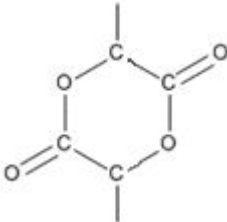
Question Number	Acceptable Answers	Reject	Mark
21(a)(iv)	<p>4-chloro-3,5-dimethylphenol</p> <p>ALLOW 3,5-dimethyl-4-chlorophenol</p> <p>ALLOW Hydroxybenzene instead of phenol</p> <p>ALLOW phen-1-ol</p> <p>IGNORE Missing / incorrect hyphens / commas / spaces</p>		(1)

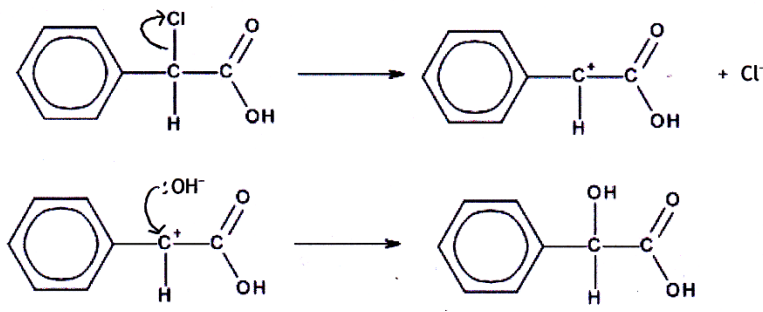
Question Number	Acceptable Answers	Reject	Mark
*21(b)	<p>Correct answer with no working scores (4)</p> <p>mol CO₂ produced = 185/24000 = 0.0077083 (1)</p> <p>mol benzoic acid = 2 x 0.0077083 = 0.015417</p> <p>ALLOW 0.01542 / 0.0154/ 0.015</p> <p>TE on mol CO₂ (1)</p> <p>mass benzoic acid = 0.015417 x 122 = 1.8808 (g)</p> <p>ALLOW 1.88124 / 1.8788 / 1.83 (g)</p> <p>TE on mol benzoic acid (1)</p> <p>% phenol = $\frac{2.5 - 1.8808}{2.5} \times 100$ = 24.767(%) = 25 (%)</p> <p>ALLOW 24.75 / 24.8 / 26.8 from earlier rounding</p> <p>TE on mass benzoic acid provided answer is <100% (1)</p> <p>IGNORE SF except 1 SF</p>		(4)

Question Number	Acceptable Answers	Reject	Mark
21(c)(i)	<p>First step</p> <p>hydrogen cyanide /HCN and potassium cyanide / KCN /cyanide ions / CN⁻ IGNORE pH in the range 5-9 / acidic medium / alkaline medium</p> <p>OR</p> <p>hydrogen cyanide / HCN and alkali / hydroxide ions / OH⁻ / pH 8-9</p> <p>OR</p> <p>potassium cyanide / KCN / cyanide ions /CN⁻ and acid / H⁺ / pH 5-6 (1)</p> <p>IGNORE ethanol / alcohol as solvent</p> <p>IGNORE heat / reflux</p> <p>Intermediate compound – stand alone</p> $ \begin{array}{c} \text{H} \quad \text{OH} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{CN} \\ \quad \\ \text{H} \quad \text{H} \end{array} $ <p>ALLOW unambiguous structural formula</p> <p>IGNORE name, even if incorrect (1)</p> <p>Second step – from correct intermediate or a compound containing CN</p> <p>IGNORE concentrations</p> <p>EITHER (Strong) acid / sulfuric acid/H₂SO₄ / hydrochloric acid / HCl /hydrogen ions / H⁺ and boil / heat / reflux</p> <p>OR Alkali / sodium hydroxide/NaOH/ potassium hydroxide /KOH/hydroxide ions /OH⁻ and boil / heat / reflux followed by (strong) acid / sulfuric acid/H₂SO₄ / hydrochloric acid / HCl / hydrogen ions / H⁺ (1)</p>	<p>Any third reagent including named acid or base</p> <p>Any third reagent</p> <p>Any third reagent</p> <p>C=N</p> <p>Alkali and acid added at same time</p>	(3)

Question Number	Acceptable Answers	Reject	Mark
21(c)(ii)	4 / four (peaks)		(1)

Question Number	Acceptable Answers	Reject	Mark
21(c)(iii)	2(.0) – 4(.0) (δ /ppm for TMS) ALLOW any number or range of numbers within the range ALLOW the range in reverse order e.g. 4(.0) – 2(.0) (δ /ppm for TMS)		(1)

Question Number	Acceptable Answers	Reject	Mark
21(c)(iv)	  Allow  ALLOW structural, displayed or skeletal formulae IGNORE additional structures as working IGNORE bond angles and bond lengths		(1)

Question Number	Acceptable Answers	Reject	Mark
21(d)(i)	 <p>ALLOW C₆H₅ for benzene ring</p> <p>First mark Curly arrow from C-Cl bond to or just beyond Cl (1)</p> <p>IGNORE dipole</p> <p>Second mark Correct intermediate and Cl⁻</p> <p>ALLOW carboxylate ion</p> <p>ALLOW Cl⁻ shown anywhere in answer (1)</p> <p>Third mark Curly arrow from O of OH⁻ to C⁺</p> <p>ALLOW the arrow to start anywhere on OH⁻, including the charge (1)</p> <p>IGNORE missing lone pair</p>	<p>Partial charge on C / Cl</p> <p>Circle missing from ring</p>	(3)

Question Number	Acceptable Answers	Reject	Mark
*21(d)(ii)	<p>First mark – stand alone A racemic mixture / racemate is formed</p> <p>OR Equal amounts / an equimolar mixture of both optical isomers / enantiomers / D-L isomers / (+) and (-) isomers (1)</p> <p>IGNORE just 'mixture is not optically active' / 'mixture does not rotate plane of plane-polarised light'</p> <p>Second mark Intermediate / carbocation is (trigonal) planar around reaction site / C⁺ / central carbon</p> <p>ALLOW Intermediate / carbocation is planar around the active site (1)</p> <p>Third mark – conditional on mention of planar (equal probability of) attacked (by nucleophile) from either side / above and below / both sides / opposite sides (of the plane) (1)</p>	<p>Carbonyl / molecule / reactant is planar</p> <p>Just 'the intermediate is planar' / the molecule is planar</p>	(3)

Question Number	Acceptable Answers	Reject	Mark
21(d)(iii)	<p>Step 1: Minimum amount of solvent to minimise the amount of mandelic acid /solid left in the solution (when it recrystallises) OR To form a saturated solution (of mandelic acid) OR To ensure that (some) acid crystallises on cooling</p> <p>ALLOW So the solution is as concentrated as possible (1)</p> <p>IGNORE Just 'to increase the yield'</p> <p>Step 2: (hot) So maximum amount / most of the solid/mandelic acid remains in (hot) solution OR To avoid the (premature) formation of crystals (in the funnel) (1)</p> <p>(filter) To remove insoluble / undissolved / solid impurities (1)</p> <p>Step 3: To ensure that maximum amount of solid crystallises</p> <p>ALLOW To obtain a better / maximum yield (of crystals)</p> <p>ALLOW So that all of the product crystallises</p> <p>IGNORE Just 'to crystallise the product' (1)</p> <p>Step 4: To remove soluble / dissolved impurities (1)</p> <p>IGNORE To speed up the process To dry the crystals</p>	<p>Speed up crystallisation</p> <p>Remove insoluble impurities</p>	(5)

(Total for Question 21 = 27 marks)

Question Number	Acceptable Answers	Reject	Mark
22(a)(i)	<p>There is (extra) stability with a full (3)d subshell / (set of 3)d orbitals / (3)d¹⁰ arrangement of electrons</p> <p>IGNORE reference to half-filled 4s orbital / repulsion in 4s²</p> <p>IGNORE just 'more stable' without some reason</p>	<p>Just 'Full (3)d orbital' / (3)d shell</p> <p>Reference to ions once only in (a)(i) and (a)(ii)</p>	(1)

Question Number	Acceptable Answers	Reject	Mark
22(a)(ii)	<p>Copper has a higher nuclear charge / more protons (so it attracts the outermost electron closer)</p> <p>IGNORE higher effective nuclear charge</p> <p>IGNORE copper has a higher charge density</p> <p>IGNORE d electrons fill an inner subshell</p> <p>IGNORE just 'stronger attraction between nucleus and (outer) electrons'</p>		(1)

Question Number	Acceptable Answers	Reject	Mark
22(b)	<p>Correct answer with no working scores (2)</p> <p>First mark – correct numbers in expression $E = 0.34 + \frac{8.31 \times 298}{9.65 \times 10^4 \times 2} \times \ln 0.100 \quad \mathbf{(1)}$</p> <p>Second mark – evaluation $= 0.34 - 0.0295$ $= (+)0.31046 / 0.3105 / 0.310 / 0.31 \text{ (V)}$</p> <p>ALLOW TE on incorrect numbers in correct formula e.g if [Cu²⁺] = 0.01 final answer is 0.28091</p> <p>No TE on incorrect formula (1)</p> <p>IGNORE SF except 1 SF</p>	(+)0.311 (V)	(2)

Question Number	Acceptable Answers	Reject	Mark
22(c)(i)	<p>First mark - E^\ominus_{cell} $E^\ominus_{\text{cell}} = 0.15 - 0.54 = -0.39 \text{ (V)}$ (1)</p> <p>Second mark – feasibility E^\ominus_{cell} is negative so reaction is not feasible</p> <p>If E^\ominus_{cell} in M1 is positive: ALLOW E^\ominus_{cell} is positive so reaction is feasible (1)</p> <p>Third mark - reason Copper(I) iodide / CuI is a solid / precipitate / ppt</p> <p>OR concentration of $\text{Cu}^+(\text{aq})$ decreases so E^\ominus for the copper half-cell increases (to more than 0.54 V and E^\ominus_{cell} becomes positive)</p> <p>ALLOW Excess iodide ions (moves equilibrium to the right) so E^\ominus for the iodine / iodide half-cell decreases (to less than 0.15 V and E^\ominus_{cell} becomes positive) (1)</p> <p>IGNORE non-standard conditions / reference to activation energy</p>		(3)

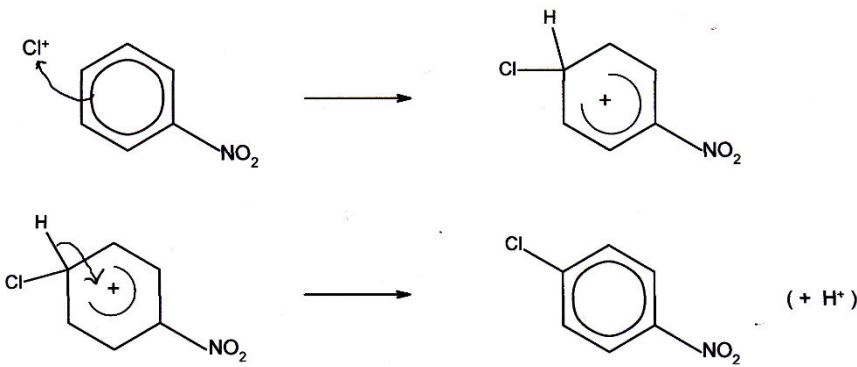
Question Number	Acceptable Answers	Reject	Mark
22(c)(ii)	<p>Correct answer to 3 SF with no working scores (4)</p> <p>mol $\text{S}_2\text{O}_3^{2-}$ used = $10.90 \times 0.150 / 1000$ = $0.001635 / 1.635 \times 10^{-3}$ (1)</p> <p>mol Cu^{2+} = $0.001635 / 1.635 \times 10^{-3}$ TE on mol $\text{S}_2\text{O}_3^{2-}$ used TE on mol I_2 if mol $\text{S}_2\text{O}_3^{2-}$ missing (1)</p> <p>EITHER mass Cu^{2+} in 25.0 cm^3 = 0.001635×63.5 = 0.10382 (g) TE on mol Cu^{2+} (1)</p> <p>mass Cu^{2+} in $1.0 \text{ dm}^3/\text{coin}$ = $0.10382 \times 1000 / 25.0$ = 4.1529 (g)</p> <p>and answer to 3 SF = 4.15 (g) TE on mass Cu^{2+} in 25.0 cm^3 (1)</p> <p>OR moles Cu^{2+} in 1.0 dm^3 = $0.001635 \times 1000 / 25.0$ = $0.0654 \text{ (mol dm}^{-3}\text{)}$ TE on mol Cu^{2+} in 25.0 cm^3 (1)</p> <p>mass Cu^{2+} in $1.0 \text{ dm}^3 / \text{coin}$ = 0.0654×63.5 = 0.10382 (g)</p> <p>and answer to 3 SF = 4.15 (g) TE on moles Cu^{2+} in 1.0 dm^3 (1)</p>		(4)

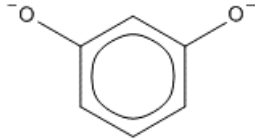
Question Number	Acceptable Answers	Reject	Mark
22(d)(i)	<p>First mark - equation</p> $\text{Ag}^{2+}(\text{aq}) + \text{Ag}(\text{s}) \rightarrow 2\text{Ag}^{+}(\text{aq}) \quad \mathbf{(1)}$ <p>ALLOW \rightleftharpoons but equation must be written in direction shown</p> <p>IGNORE missing / incorrect state symbols</p> <p>Second mark – explanation, conditional on M1</p> <p>No, this is the reverse of disproportionation / comproportionation / Ag^{+} is oxidised and reduced in the reverse reaction</p> <p>OR</p> <p>No, it must be an element in a single species that is both oxidised and reduced / 2 different species are oxidised and reduced</p> <p>OR</p> <p>2 different oxidation states are not produced</p> <p>ALLOW</p> <p>No, as only 1 species is produced</p> <p>No TE on incorrect equation (1)</p> <p>IGNORE just 'not disproportionation'</p>		(2)

Question Number	Acceptable Answers	Reject	Mark
22(d)(ii)	<p>ALLOW oxidation numbers written by correct species in equation</p> <p>EITHER</p> <p>Au: 0 to (+)3 and oxidation (1)</p> <p>N: (+)5 to (+)2 and reduction (1)</p> <p>OR</p> <p>All oxidation numbers</p> <p>Au: 0 and (+)3</p> <p>N: (+)5 and (+)2 (1)</p> <p>Au is oxidised and N is reduced (1)</p> <p>No TE on incorrect oxidation numbers</p> <p>ALLOW oxidation numbers as Roman numerals / 3+, 2+, 5+ and as charges e.g. Au³⁺</p> <p>IGNORE oxidation numbers of other elements</p>		(2)

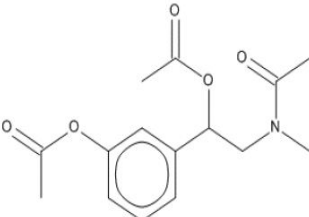
(Total for Question 22 = 15 marks)
Total for Section B = 51 MARKS

Section C

Question Number	Acceptable Answers	Reject	Mark
23(a)(i)	<p>First mark - formation of electrophile $\text{Cl}_2 + \text{AlCl}_3 \rightarrow \text{Cl}^+ + \text{AlCl}_4^- / [\text{AlCl}_4]^- / \delta^+\text{Cl}-\text{AlCl}_4 \delta^-$ (1)</p> <p>Mechanism Note - If benzene used instead of nitrobenzene / if final product is not 1-chloro-4-nitrobenzene, do not award the mark for the intermediate</p>  <p>Second mark Curly arrow from on or within the circle to Cl^+</p> <p>ALLOW Curly arrow from anywhere within the hexagon</p> <p>ALLOW Curly arrow to any part of the Cl^+, including to the + charge</p> <p>ALLOW Cl with no charge if M1 not awarded, but do not allow any other electrophile (1)</p> <p>Third mark Intermediate structure including charge with horseshoe covering at least 3 carbon atoms and facing the tetrahedral carbon and some part of the positive charge must be within the horseshoe</p> <p>ALLOW dashed / dotted line for horseshoe (1)</p> <p>Fourth mark Curly arrow from C-H bond to anywhere in the hexagon reforming the delocalised structure / (1)</p> <p>Correct Kekulé structures score full marks</p> <p>IGNORE any involvement of AlCl_4^- in the final step</p>	<p>Any Friedel-Crafts catalyst except AlCl_3</p> <p>Curly arrow on or outside the hexagon</p> <p>Dotted bonds to H and Cl unless clearly part of a 3D structure</p>	<p>(4)</p>

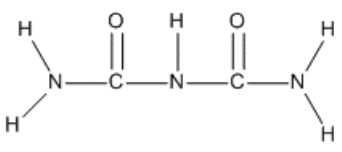
Question Number	Acceptable Answers	Reject	Mark
23(a)(ii)	 <p>ALLOW O⁻ Na⁺ / ONa</p> <p>ALLOW OH</p> <p>ALLOW Kekulé structure</p>	<p>O-Na C-NaO</p> <p>OH-C C-HO</p>	(1)

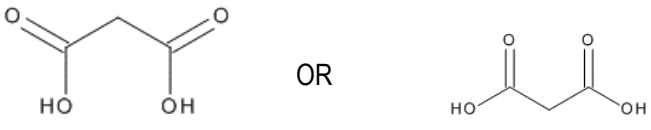
Question Number	Acceptable Answers	Reject	Mark
23(a)(iii)	<p>C₁₂H₉N₃O₄</p> <p>ALLOW symbols in any order e.g. C₁₂H₉O₄N₃</p> <p>IGNORE any working before the formula</p>		(1)

Question Number	Acceptable Answers	Reject	Mark
23(b)	 <p>First mark Ethanoylation of NH group (1)</p> <p>Second mark Ethanoylation of both OH groups</p> <p>ALLOW Ethanoylation of alkyl OH group only or phenol group only (1)</p> <p>ALLOW Structural, displayed or skeletal formula or any combination</p>		(2)

Question Number	Acceptable Answers	Reject	Mark
23(c)(i)	<p data-bbox="339 206 671 239">Examples of diagrams</p> <div data-bbox="363 245 925 657"> </div> <p data-bbox="339 672 1058 814">First mark – stand alone Hydrogen bonds can form between urea and water - this may be labelled on the diagram (1)</p> <p data-bbox="339 821 957 886">IGNORE Hydrogen bonds between urea molecules</p> <p data-bbox="339 924 1023 1026">Second mark – position of hydrogen bond Between the O–H in water and N–H /O=C in urea</p> <p data-bbox="339 1069 1050 1138">ALLOW any of the three positions described or shown in a diagram (1)</p> <p data-bbox="339 1177 957 1278">IGNORE bond angle in diagram Hydrogen bonds between urea molecules</p> <p data-bbox="339 1317 1023 1386">Third mark – linear hydrogen bond Linear N–H····O /O–H····O / N····H–O bond</p> <p data-bbox="339 1425 847 1457">ALLOW bond angle stated as 180°</p> <p data-bbox="339 1496 962 1565">ALLOW this mark if one hydrogen bond is shown linear</p> <p data-bbox="339 1604 1058 1673">ALLOW 180° bond angle in hydrogen bond between urea molecules (1)</p> <p data-bbox="339 1711 1038 1813">Note Full marks can only be awarded if a diagram is shown</p>		(3)

Question Number	Acceptable Answers	Reject	Mark
23(c)(ii)	$2\text{NH}_3 + \text{CO}_2 \rightarrow (\text{H}_2\text{N})_2\text{CO} + \text{H}_2\text{O}$ ALLOW molecular formula or other correct structural formula for urea e.g. CON_2H_4 , NH_2CONH_2 , $\text{CO}(\text{NH}_2)_2$ ALLOW multiples IGNORE state symbols, even if incorrect		(1)

Question Number	Acceptable Answers	Reject	Mark
23(c)(iii)	 ALLOW Terminal NH_2 and central NH IGNORE bond lengths and angles IGNORE structural formula – $\text{NH}_2\text{CONHCONH}_2$	$\text{NH}_2\text{-C}$	(1)

Question Number	Acceptable Answers	Reject	Mark
23(c)(iv)	 IGNORE connectivity of OH IGNORE additional formulae as working	Structural or displayed formulae	(1)

Question Number	Acceptable Answers	Reject	Mark
23(d)(i)	Alkene, nitrile and ester All 3 scores 2 Any 2 scores 1 IGNORE alkyl / alkane IGNORE Cyanide / cyano IGNORE formulae eg $\text{C}=\text{C}$		(2)

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