



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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**CHEMISTRY**

**0620/32**

Paper 3 Theory (Core)

**October/November 2017**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

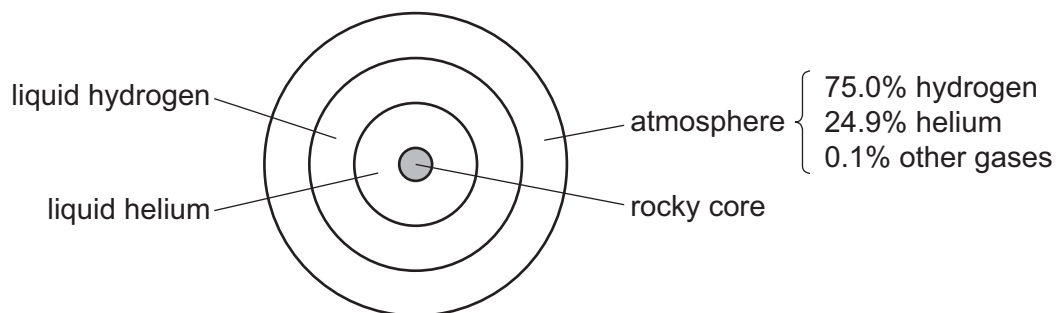
The number of marks is given in brackets [ ] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **15** printed pages and **1** blank page.



2 The diagram shows the composition of the planet Saturn.



(a) Describe how Saturn's atmosphere differs from the Earth's atmosphere.  
Give **three** differences.

- 1 .....
- .....
- 2 .....
- .....
- 3 .....
- .....

[3]

(b) Some properties of hydrogen and helium are given in the table.

element	density of the liquid in g/cm <sup>3</sup>	melting point in °C	boiling point in °C
hydrogen	0.07	-259	-253
helium	0.15	-272	-269

(i) Use the information to suggest why the layer of liquid hydrogen in Saturn floats on top of the liquid helium.

..... [1]

(ii) What is the physical state of hydrogen at -250 °C?  
Explain your answer.

.....

..... [2]

(c) The atmosphere of Saturn contains small amounts of ammonia.

(i) Describe a test for ammonia.

test .....

result .....

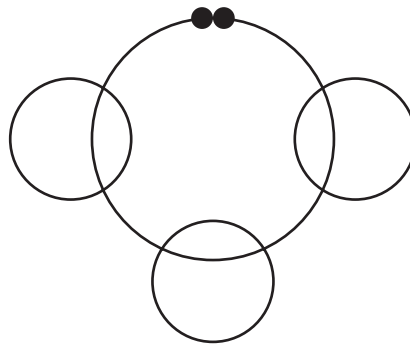
[2]

(ii) Ammonia is a covalent compound.

Complete the diagram to show

- the arrangement of electrons in a molecule of ammonia,
- the symbols of the atoms present.

Show outer electrons only.



[2]

(d) Saturn's atmosphere also contains small amounts of ammonium hydrosulfide.

Calculate the relative molecular mass of ammonium hydrosulfide,  $\text{NH}_4\text{SH}$ .  
Use your Periodic Table to help you.

relative molecular mass = ..... [2]

(e) Saturn's atmosphere also contains small amounts of methane.

Small amounts of methane are present in the Earth's atmosphere.  
Methane is a greenhouse gas.

(i) Name another greenhouse gas present in the Earth's atmosphere.

..... [1]

(ii) Scientists are concerned about the increase in the amount of greenhouse gases in the Earth's atmosphere.

Explain why.

..... [1]

[Total: 14]

3 The following compounds are present in a liquid used for cleaning metal.

ethanoic acid  
 ethanol  
 glycerol  
 sodium chloride  
 water

(a) (i) Draw the structure of the functional group present in ethanoic acid.

[1]

(ii) Which **one** of the following pH values is acidic?  
 Put a circle around the correct answer.

pH4          pH7          pH9          pH13

[1]

(iii) Ethanoic acid reacts with sodium hydroxide.

What type of reaction is this?  
 Put a circle around the correct answer.

**neutralisation**          **oxidation**          **polymerisation**          **reduction**

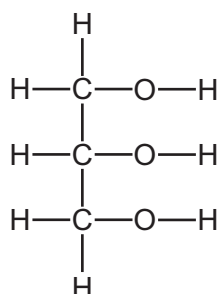
[1]

(iv) The reaction of ethanoic acid with sodium hydroxide is exothermic.

What is meant by the term *exothermic*?

..... [1]

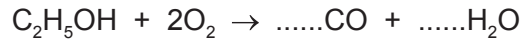
(b) The structure of glycerol is shown.



Deduce the molecular formula of glycerol showing the number of carbon, hydrogen and oxygen atoms.

..... [1]

(c) Balance the chemical equation for the incomplete combustion of ethanol.



[2]

(d) (i) Describe a method of obtaining pure samples of **both** sodium chloride and water from aqueous sodium chloride. Explain why this method works.

.....

.....

.....

.....

..... [3]

(ii) Which physical property could you measure to find out if a sample of water is pure?

..... [1]

(iii) Sodium chloride contains chloride ions.

Describe a test for chloride ions.

test .....

result .....

[2]

[Total: 13]

4 The table shows the properties of four substances.

substance	boiling point	electrical conductivity of solid	electrical conductivity when molten	solubility in water
calcium iodide	very high	does not conduct	conducts	
phosphorus	low	does not conduct	does not conduct	insoluble
sodium chloride	very high	does not conduct	conducts	soluble
zinc	high		conducts	insoluble

(a) Complete the table to show the solubility in water of calcium iodide and the electrical conductivity of solid zinc. [2]

(b) Give **one** piece of evidence from the table that shows that phosphorus is a simple covalent substance.

..... [1]

(c) What information in the table shows that sodium chloride is an ionic compound?

.....  
 ..... [2]

(d) Molten calcium iodide can be electrolysed.

Predict the products of this electrolysis at

the positive electrode (anode), .....

the negative electrode (cathode). .....

[2]

(e) An atom of phosphorus has 31 nucleons.

Deduce the number of protons and neutrons present in **one** atom of phosphorus.  
 Use your Periodic Table to help you.

number of protons .....

number of neutrons .....

[2]

(f) Phosphorus burns in an excess of air to form phosphorus(V) oxide.

Is phosphorus(V) oxide an acidic oxide or a basic oxide?  
Explain your answer.

.....  
..... [1]

(g) Arsenic is in the same group of the Periodic Table as phosphorus.  
Arsenic sublimes at 613 °C.

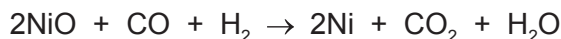
What is meant by the term *sublimation*?

.....  
..... [1]

[Total: 11]



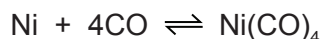
- 5 Nickel can be obtained from nickel(II) oxide by heating it with a mixture of carbon monoxide and hydrogen.



- (a) How does this equation show that the nickel(II) oxide is reduced?

..... [1]

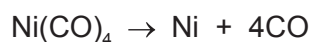
- (b) Nickel can be purified by reacting impure nickel with carbon monoxide. A compound called tetracarbonylnickel,  $\text{Ni}(\text{CO})_4$ , is formed.



What is the meaning of the symbol  $\rightleftharpoons$ ?

..... [1]

- (c) The tetracarbonylnickel is heated to obtain pure nickel.



- (i) Suggest why the nickel obtained can be separated easily from the carbon monoxide.

..... [1]

- (ii) State **one** adverse effect of carbon monoxide on health.

..... [1]

- (d) Nickel is a transition element.  
Potassium is a Group I element.

- (i) Describe **two** differences in the physical properties of nickel and potassium.

1 .....

.....

2 .....

.....

[2]

- (ii) Describe **one** difference in the properties of nickel(II) chloride and potassium chloride.

..... [1]

- (e) The properties and relative reactivity with water of some Group I elements are shown in the table.

element	density in g/cm <sup>3</sup>	boiling point in °C	relative reactivity with water
sodium		883	forms bubbles rapidly but does not burst into flames
potassium	0.86	760	forms bubbles very rapidly and bursts into flames
rubidium	1.53		
caesium	1.88	669	reacts explosively

- (i) Complete the table

- to predict the boiling point of rubidium,
- for the relative reactivity of rubidium with water.

[2]

- (ii) Describe the general trend in the density of the Group I elements.

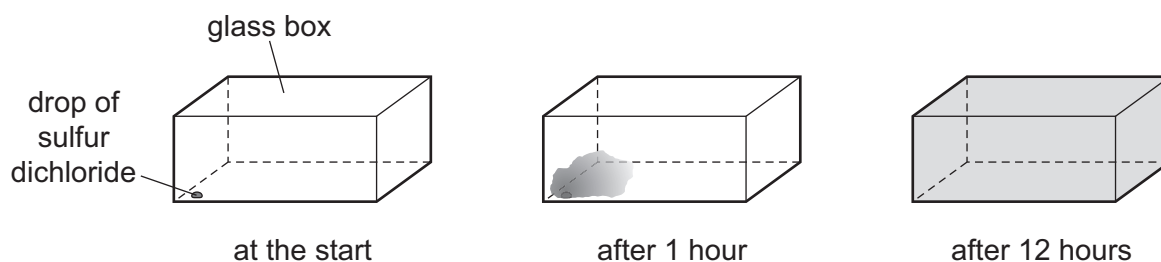
..... [1]

[Total: 10]



7 Sulfur dichloride,  $\text{SCl}_2$ , is a red liquid which vaporises easily at room temperature and pressure.

- (a) A drop of sulfur dichloride was placed in the corner of a glass box. The glass box was closed and left for 12 hours. After 12 hours a red vapour had spread to fill the whole box.



Explain these observations using the kinetic particle model.

.....

.....

.....

.....

.....

..... [3]

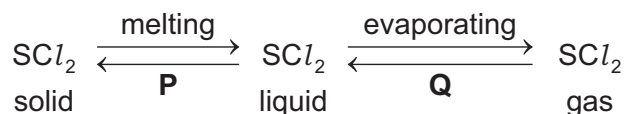
- (b) Sulfur dichloride can be made by passing chlorine through liquid disulfur dichloride,  $\text{S}_2\text{Cl}_2$ .

Complete the chemical equation for this reaction.



[2]

- (c) Some changes of state of sulfur dichloride are shown.



Name the changes of state represented by **P** and **Q**.

**P** .....

**Q** .....

[2]

[Total: 7]

8 Calcium carbonate (limestone) decomposes when heated.

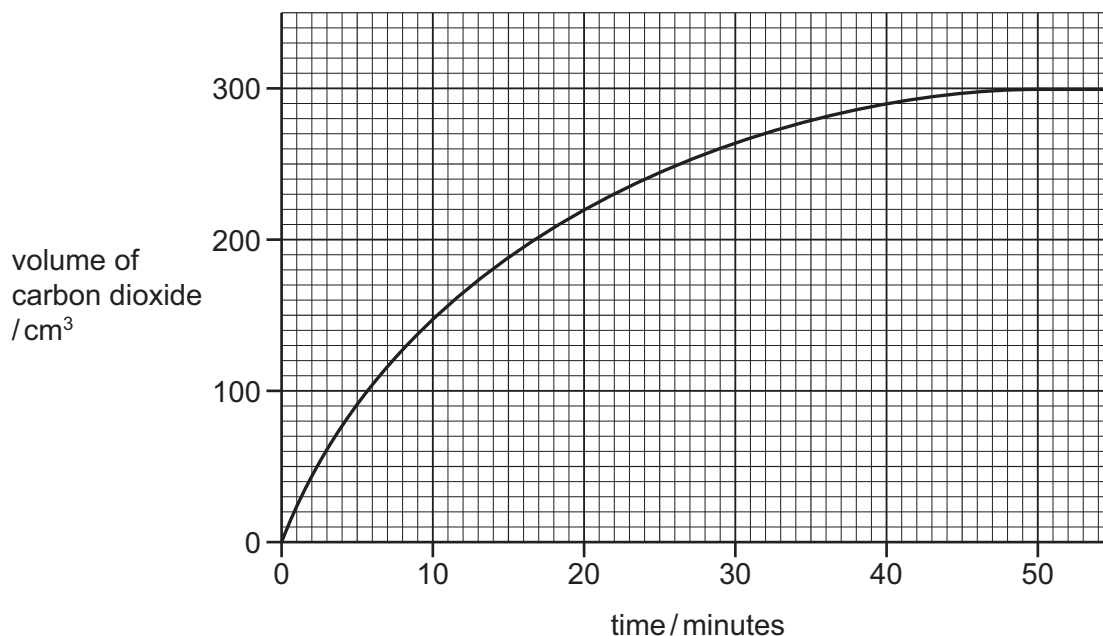


(a) When 20.0g of calcium carbonate are decomposed, 11.2g of calcium oxide (lime), CaO, are formed.

Calculate the mass of calcium oxide formed when 160.0g of calcium carbonate are decomposed.

..... g [1]

(b) The graph shows the volume of carbon dioxide produced when some small pieces of calcium carbonate are heated and decompose.



(i) Deduce the volume of carbon dioxide produced during the first 20 minutes of the decomposition.

..... [1]

(ii) At what time was the reaction complete?

..... [1]

(iii) What would be the effect, if any, on the rate of reaction if the same mass of powdered calcium carbonate were used?

..... [1]

(c) The table shows how limestone is used.

use of limestone	percentage of limestone used for this purpose
agriculture	
cement manufacture	37
chemical industries	14
iron and steel manufacture	11
road building	20
other uses	2
total	100

(i) What percentage of limestone is used in agriculture?

..... [1]

(ii) Limestone and lime are used in agriculture.

Why is lime used in agriculture?

.....  
 ..... [2]

[Total: 7]

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