
CHEMISTRY

0620/63

Paper 6 Alternative to Practical

October/November 2017

MARK SCHEME

Maximum Mark: 40

Published

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This document consists of **5** printed pages.

Question	Answer	Marks
1(a)	pestle	1
	(teat) pipette	1
1(b)	to increase surface area / make it dissolve faster	1
1(c)	nitric (acid)	1
1(d)	residue	1
1(e)	M1 add a more reactive metal (e.g. zinc / magnesium)	1
	M2 displaces lead / filter out lead	1

Question	Answer	Marks
2(a)	temperature boxes completed: 23, 16, 14, 13, 12, 11, 11, 11, 11, 11 all readings correct = [2] 8 or 9 readings correct = [1]	2
2(b)	temperature boxes completed correctly: 22, 26, 29, 31, 32, 33, 34, 35, 35, 35 all readings correct = [2] 8 or 9 readings correct = [1]	2
2(c)	all points plotted	1
	two smooth line graphs (one line graph correct = [1])	2
	both graphs appropriately labelled	1
2(d)(i)	value from graph	1
	shown clearly	1

Question	Answer	Marks														
2(d)(ii)	value from graph	1														
	shown clearly	1														
2(e)	exothermic	1														
2(f)	room temperature / 22 °C AND reaction has finished / all the solid has dissolved	1														
2(g)	<table border="1"> <thead> <tr> <th data-bbox="407 545 1052 596">source of error</th> <th data-bbox="1052 545 1854 596">improvement</th> </tr> </thead> <tbody> <tr> <td data-bbox="407 596 1052 647">heat losses</td> <td data-bbox="1052 596 1854 647">use a lid / lag the apparatus</td> </tr> <tr> <td data-bbox="407 647 1052 699">use of a measuring cylinder</td> <td data-bbox="1052 647 1854 699">use a pipette/burette</td> </tr> <tr> <td data-bbox="407 699 1052 750">wet cup in the second experiment</td> <td data-bbox="1052 699 1854 750">use new/another cup OR dry the cup</td> </tr> <tr> <td data-bbox="407 750 1052 801">the solid absorbs water from the air</td> <td data-bbox="1052 750 1854 801">store in a sealed container / airtight container / desiccator</td> </tr> <tr> <td data-bbox="407 801 1052 852">only done once</td> <td data-bbox="1052 801 1854 852">repeat and average</td> </tr> <tr> <td data-bbox="407 852 1052 935">different masses of solids used / masses of solids not measured</td> <td data-bbox="1052 852 1854 935">use same mass of solid / weigh the solids</td> </tr> </tbody> </table>	source of error	improvement	heat losses	use a lid / lag the apparatus	use of a measuring cylinder	use a pipette/burette	wet cup in the second experiment	use new/another cup OR dry the cup	the solid absorbs water from the air	store in a sealed container / airtight container / desiccator	only done once	repeat and average	different masses of solids used / masses of solids not measured	use same mass of solid / weigh the solids	4
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2(h)	fewer data / less detail / fewer readings / graph not as good / not enough readings whilst the solid is reacting	1														

Question	Answer	Marks
3(a)(i)	green	1
	precipitate	1
3(a)(ii)	green solution / precipitate dissolves	1
3(a)(iii)	bubbles / fizzing / effervescence	1
	(red) litmus paper / Universal Indicator paper	1
	(red litmus paper) turns blue / (Universal Indicator paper) turns purple	1
3(b)	ammonia / NH ₃	1
3(c)	(aqueous) ammonia / NH ₃	1

Question	Answer	Marks
4	<p><i>heating to dryness method</i></p> <p>max [6]: M1 weigh (any) sample of washing soda M2 heat (to remove water of crystallisation) M3 in named container M4 cool M5 reweigh M6 repeat heating M7 to constant mass M8 appropriate calculation suggested for the percentage of water</p> <p><i>mass of water method</i></p> <p>max [6]: M1 weigh (any) sample of washing soda M2 heat to remove water of crystallisation M3 in named container M4 using apparatus capable of collecting water (vapour) M5 cool / condense (water vapour) M6 continue until no more collects M7 weigh water M8 appropriate calculation suggested for the percentage of water</p>	6